

Effect of the electrodeposition potential on the magnetic properties of FeCoNi Ims

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and potential are essential variables that should be adjusted to control kinetics of the deposition process. By altering the electrodeposition potential, microstructure and chemical composition of the transition metals alloys, such as CoNiCu [15] and CoNi [16], can be modified. This condition leads to significant changes in their magnetic properties [10, 16]. Since the magnetic properties of the FeCoNi also depend on its microstructure and composition, it is interesting to investigate the role of these electrodeposition variables on the electrodeposited FeCoNi films.

In our previous works [15, 16], the FeCoNi film was successfully electrodeposited on a flexible substrate of polyethylene terephthalate coated with indium tin oxide (ITO-PET). The using of a flexible substrate was motivated by the recent demand for flexible magnetic films due to the development of flexible electronic devices [17, 18]. Herein, we investigated the influence of electrodeposition potential on magnetic properties of the FeCoNi film electrodeposited on the ITO-PET substrate. The FeCoNi deposits obtained in this work demonstrated properties of typical soft magnetic alloys with low coercivity and high saturation which could be controlled through the applied electrodeposition potential. The relation between the applied potential, chemical composition and magnetic properties of the film was discussed.

2. Materials and method

The electrodeposition of FeCoNi was performed in sulfate electrolyte system, which was freshly prepared from analytical grade FeSO₄·7H₂O, CoSO₄·7H₂O, NiSO₄·6H₂O, H₃BO₃ precursors and the food grade C₇H₄NNaO₃S·2H₂O chemical. The concentration of the electrolytes used in the electrodeposition process for the FeCoNi preparation is listed in Table 1. The electrodeposition was carried out under various electrodeposition potentials using EDAQ potentiostat of EA163 type in the potentiostatic mode. The substrate used in this work was polyethylene terephthalate coated with indium tin oxide (ITO-PET) with a sheet resistance of 10 Ω/sq supplied by Kintec. The coated ITO is amorphous

Table 1. Concentration of the solution used as metal sources in the electrodeposition process.

Chemicals	Molarities [mol/L]
FeSO ₄ ·7H ₂ O	0.010
CoSO ₄ ·7H ₂ O	0.020
NiSO ₄ ·6H ₂ O	0.170
H ₃ BO ₃	0.400
C ₇ H ₄ NNaO ₃ S·2H ₂ O	0.008

phase composed of In₂O₃:SnO₂ at a ratio of 9:1 with a thickness of 180 nm. A platinum wire and Ag/AgCl acted as a counter and a reference electrode, respectively. The electrodeposition potential applied in the process was varied from 1.00 V to 2.00 V relative to Ag/AgCl electrode with the deposition time of 5 minutes. In this case, the electrodeposition processes were carried at potentials higher than the standard reduction potential of Fe²⁺ (0.44 V), Co²⁺ (0.28 V), and Ni²⁺ (0.25 V) due to the co-deposition overpotential of the metal ions [19, 20].

The mean crystallite size of FeCoNi films was evaluated by Panalytical EMPYREAN X-ray diffractometer (XRD) employing CuK_α radiation (λ = 1.541874 Å). The obtained XRD patterns were compared to the standard pattern of FeCoNi alloy from Crystallography Open Database (COD) number 96-900-00900 [21]. The mean crystallite size of the FeCoNi films was then determined by Rietveld refinement method using HighScore Plus software. The mass content of Fe, Co and Ni in the deposits was quantitatively determined using a Shimadzu AA7000 atomic absorption spectrometer (AAS). Magnetic properties were measured for 4 mm × 4 mm FeCoNi films using LakeShore 7404 vibrating sample magnetometer (VSM) equipped with a magnetic field (H) up to 900 kA/m.

3. Results and discussion

X-ray diffraction patterns of the FeCoNi alloys deposited at various negative potentials are compared in Fig. 1. Obviously, all patterns show a diffraction peak at a diffraction angle 2θ of 44.17° related to (1 1 1) plane

of FeCoNi. The sample prepared at a low potential (1.00 V vs. Ag/AgCl) shows additional peaks corresponding to the diffraction peaks of ITO and PET substrates [22]. Since all of the FeCoNi Im were electrodeposited in the same deposition time, the pattern was then attributed to very thin deposits due to the low deposition rate of Fe, Co and Ni.

At 1.00 V vs. Ag/AgCl and hence, it allowed the X-ray to reach the substrate. Nevertheless, these additional peaks vanished in the Im obtained at more negative values of the electrodeposition potential (1.10 V to 2.00 V vs. Ag/AgCl). It is also noted that another diffraction peak appears at a different diffraction angle 51.57 which corresponds to the (2 0 0) plane of FeCoNi [21]. This peak was observed in all samples prepared through electrodeposition at potential ranging from 1.50 V to 2.00 V vs. Ag/AgCl. All diffraction patterns shown in Fig. 1 have been matched with the results on the phase structure evaluation of mechanically alloyed FeCoNi reported by Pikula [3]. The formation of FCC solid solution of FeCoNi alloy in the Im samples obtained in the current work was then confirmed. Additionally, the patterns shown in Fig. 1 are characterized by broadened peaks reflecting the existence of ultra fine crystallites in the Im. The mean crystallite size was found in a small range between 11 nm and 15 nm. Hence, there is no substantial influence of electrodeposition potential on the mean crystallite size of the FeCoNi Im. However, when the mean crystallite size of FeCoNi Im is compared with that obtained by mechanical alloying technique [3], it is realized that the mean crystallite size of the current alloy Im is much smaller than that of mechanically alloyed FeCoNi.

The atomic fraction (at.%) of Fe, Co and Ni contained in the FeCoNi alloy Im deposited at different electrodeposition potentials is plotted in Fig. 2a. The plot indicates that the electrodeposition potentials applied during the Im deposition can control the atomic fraction of Fe, Co and Ni atoms. The more negative potential is applied, the more abundant concentration of Ni is in the Im. In addition, it can be seen that Fe and Co concentrations were found to decrease at more negative potentials. In this case, Fe content decreased from 22.68 at.%

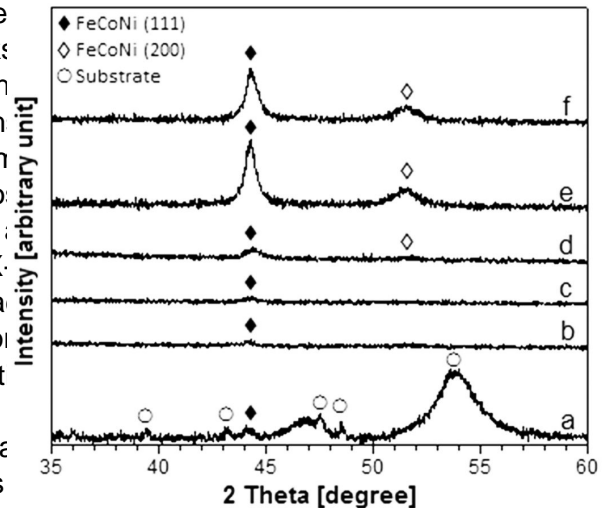


Fig. 1. X-ray diffraction patterns of the FeCoNi Im deposited at electrodeposition potentials of (a) 1.00 V, (b) 1.10 V, (c) 1.30 V, (d) 1.50 V, (e) 1.75 V and (f) 2.00 V vs. Ag/AgCl.

at a potential of 1.10 V vs. Ag/AgCl to 10.18 at.% at a potential of 2.00 V vs. Ag/AgCl. Similarly, the Co content decreased from 40.77 at.% to 20.70 at.%. Consequently, the alloy was enriched with Ni as the applied potential was getting more negative. Fig. 2b shows the plots of Ni/Fe and Ni/Co atomic ratios versus the applied potential. As the potential increases negatively from 1.00 V to 1.75 V vs. Ag/AgCl, the Ni/Fe and Ni/Co ratios increase almost linearly. Nevertheless, at a potential lower than 1.75 V vs. Ag/AgCl, no further increase is noticed. The difference between the ratio values of Ni/Fe and Ni/Co must be related to the standard potential of Ni^{2+} .

The relatively high content of Co and Fe in the Im obtained from electrolyte with high Ni^{2+} concentration, especially observed at 0 V vs. Ag/AgCl, indicates anomalous co-deposition in the electrodeposition process. Anomalous co-deposition is a phenomenon showing a high deposition preference of less noble metal, namely Fe and Co, compared to more noble metal Ni. However, at more negative potentials (1.10 V vs. Ag/AgCl) the Ni content was found to increase gradually when the electrodeposition potential shifted to more negative ones. It can

be assumed that the anomalous co-deposition characteristics decreased at more negative potentials. This results could be attributed to the high overpotential of Ni [19, 23] which leads to rising the deposition rate of Ni at more negative potential. It was also found that the change of Co content at different electrodeposition potentials is similar to that of Fe (Fig. 2a). In fact, the standard reduction potential of Co is closer to that of Ni rather than Fe. This case can be related to the composition ratio value (CRV) of Co that exhibits similar behavior to that of Fe [24].

It was recorded in our previous report that although the standard reduction potential of Ni^{2+} is lower than that of both Fe^{2+} and Co^{2+} , the Ni^{2+} ion possesses the highest overpotential among the investigated metal ions [14]. The high overpotential of Ni^{2+} has led to a higher deposition rate of Ni at a high electrodeposition potential difference compared with those of Fe and Co. Consequently, the atomic fraction of Ni in the film is much higher than that of Fe and Co. In Fig. 2b, the atomic ratio between Co and Fe was also plotted. There is almost no change in the atomic ratio between Co and Fe atoms with the applied potential in the range of 1.1 V to 2.00 V vs. Ag/AgCl. However, at a potential of 1.00 V vs. Ag/AgCl, the atomic fractions of the three components are mostly equal to each other with the composition ratio about 1:1:1.

Fig. 3a compares the hysteresis loops of the FeCoNi films prepared at potentials of 1.00 V, 1.30 V and 1.75 V. Evaluation of magnetic properties of FeCoNi alloy films was carried out in two different directions. The two directions were of each sample is the same for both directions (IP marked as IP (in-plane) and OP (out-plane) which refer to parallel and perpendicular to the direction of the applied magnetic field, respectively. Hysteresis loops shown in Fig. 3a are typical of that of soft magnetic films. The saturation magnetization is achieved mostly at a relatively small external magnetic field. This result is confirmed if the preferred orientation of the crystallites of magnetic phase are oriented to a preferred direction. For this case, it still cannot be confirmed if the preferred orientation of the crystallites can induce the magnetic anisotropy of the FeCoNi films prepared at the electrodeposition potentials of 1.00 V, 1.30 V and 1.75 V vs. Ag/AgCl. For the latter, a slightly higher magnetic field was required to achieve the saturation state due to the anisotropy characteristic of the sample. Although the sample shows anisotropy by the sample prepared at an electrodeposition

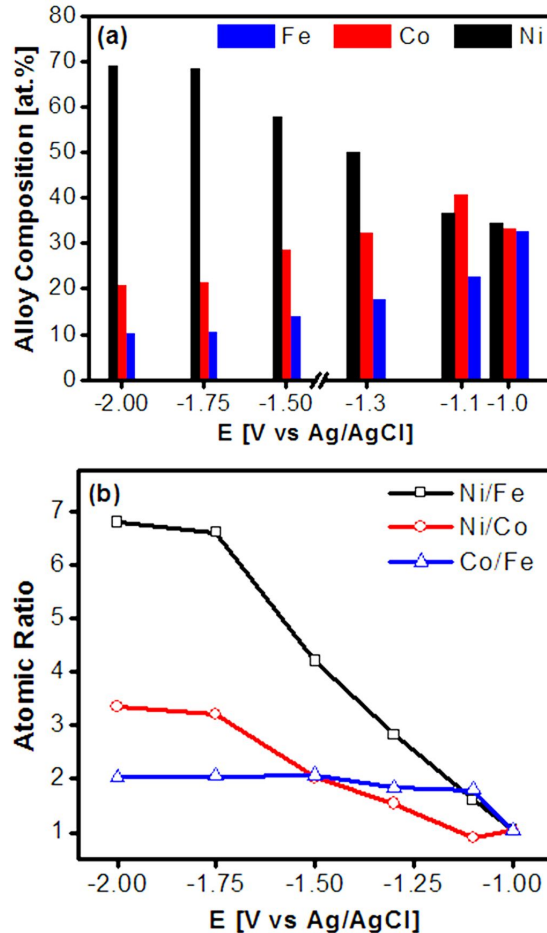


Fig. 2. Atomic composition of Ni, Co and Fe (a) and atomic ratio Ni/Fe, Ni/Co, Co/Fe values (b) in FeCoNi films prepared at different electrodeposition potentials.

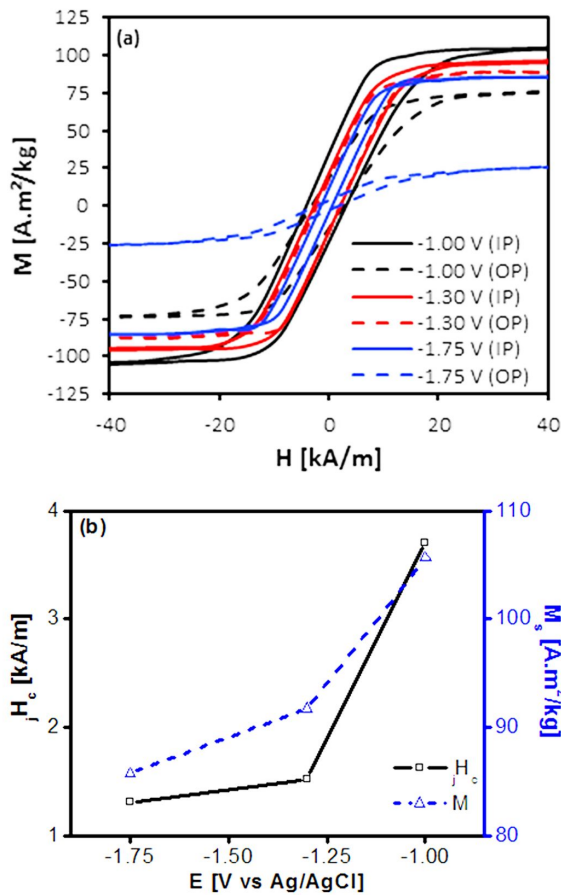


Fig. 3. Hysteresis loops of FeCoNi Ims (a) prepared at various electrodeposition potentials and (b) the corresponding intrinsic coercivity (jH_c) and saturation magnetization (M_s) values obtained from IP loops.

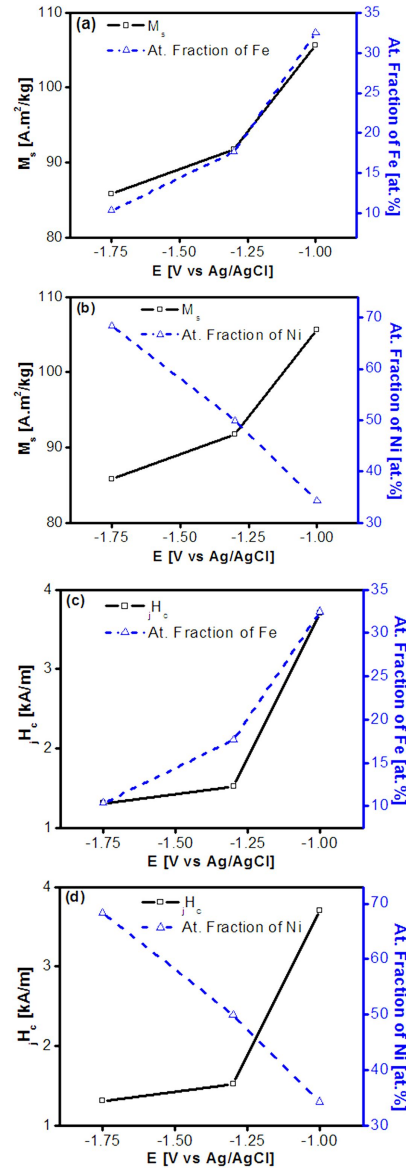


Fig. 4. Composition dependent intrinsic coercivity (jH_c) and saturation magnetization (M_s) values of FeCoNi Ims measured in IP directions

potential of -1.75 V vs. Ag/AgCl is more likely due to a shape anisotropy of the sample.

Both saturation magnetization and coercivity values of the samples increased as the electrodeposition potential became less negative. The increase in saturation magnetization value is due to the higher content of Fe and Co atoms, and lower content of Ni in the FeCoNi alloy. This condition was followed by an increase in the corresponding coercivity.

In the previous discussion, we have shown that the applied potential during electrodeposition of FeCoNi Ims caused a change in the composition of the alloy (Fig. 4b and 4d). The reason for these results is that the net magnetic moment of Fe atom ($2\mu_B$) [25]

resulted in a modification in the magnetic properties of FeCoNi Ims. The plots of M_s and jH_c values of FeCoNi Ims of various compositions are shown in Fig. 4. It shows an increase in both M_s and jH_c values with an increase in Fe content in the FeCoNi Im (Fig. 4a and 4c). The M_s and jH_c values get increased with a decrease in Ni content (Fig. 4b and 4d). The reason for these results is that

is the highest among the three transition metals followed by Co (1.72 μ B) and Ni (0.60 μ B) [25]. It is known that magnetization and coercivity are properties which depend on microstructure and chemical composition of materials [5, 10, 26]. However, since there is no meaningful change in the crystallite size of the electrodeposited FeCoNi alloys, the changes in their magnetic characteristics could be solely attributed to the chemical composition of the alloy. In this case, once again the increase in Fe and Co content resulted in the increasing saturation magnetization. However, the change in coercivity is more likely due to the change in the magnetic crystalline constant of the FeCoNi.

4. Conclusion

This study demonstrates the influence of the applied potential on magnetic properties of electrodeposited FeCoNi films. The magnetic properties were found to be solely dependent on the atomic fraction of Fe, Co and Ni atoms, which was successfully controlled through the adjustment of the electrodeposition potential because there was no significant change found in the crystallite size. The FeCoNi films possess excellent soft magnetic properties with the highest M value of 105 Am²/kg and jH_c value of 3.7 kA/m obtained at the electrodeposition potential of 1.00 V vs. Ag/AgCl with the Fe, Co and Ni composition close to 1:1:1.

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