

A. Kausar**National University of Sciences and Technology, Islamabad, Pakistan***asheesgreat@yahoo.com*

STRUCTURE AND PROPERTIES OF POLYACRYLONITRILE/ POLYSTYRENE AND CARBON NANOPARTICLE-BASED NANOCOMPOSITE FOAMS

ABSTRACT

In this study, novel polyacrylonitrile/polystyrene (PAN/PS) blend has been prepared and reinforced with carbon nanoparticle to form polyacrylonitrile/polystyrene/carbon nanoparticle (PAN/PS/CNP) nanocomposite foam. Acid-functional carbon nanoparticle (0.1-3 wt.%) was used as nano-reinforcement for PAN/PS blend matrix. 2'-azobisisobutyronitrile was employed as foaming agent. The PAN/PS/CNP nanocomposite foams have been tested for structure, morphology, mechanical properties, thermal stability, non-flammability, water uptake, and toxic ion removal. Field-emission scanning electron microscopy and transmission electron microscopy exposed unique nanocellular morphology owing to physical interaction between the matrix and functional CNP. PAN/PS/CNP 0.1 Foam with 0.1 wt.% nanofiller had compression strength, modulus, and foam density of 41.8 MPa, 22.3 GPa, and 0.9 mgcm⁻³, respectively. Nanofiller loading of 3wt.% (PAN/PS/CNP 3 Foam) considerably enhanced the compression strength, modulus, and foam density as 68.2 MPa, 37.7 GPa, and 1.9 mgcm⁻³, respectively. CNP reinforcement also enhanced the initial weight loss and maximum decomposition temperature of PAN/PS/CNP 3 Foam to 541 and 574 °C, relative to neat foam ($T_0 = 411$ °C; $T_{10} = 459$ °C). Nanocomposite foams have also shown excellent flame retardancy as V-0 rating and high char yield of up to 57% were attained. Due to hydrophilic nature of functional carbon nanoparticle, water absorption capacity of 3 wt.% nanocomposite foam was 30% higher than that of pristine foam. Moreover, novel foams were also tested for the removal of toxic Pb²⁺ ions. PAN/PS/CNP 3 Foam has shown much higher ion removal capacity (166 mg/g) and efficiency (99 %) than that of PAN/PS foam having removal capacity and efficiency of 90 mg/g and 45 %, respectively.

Keywords: polyacrylonitrile; polystyrene CNP; foam; Pb²⁺ ions

INTRODUCTION

Polyacrylonitrile (PAN) is an acrylic thermoplastic polymer having range of useful properties such as high strength, modulus, and thermal stability [1]. PAN occur as film, fiber, foam, and other useful polymeric forms [2, 3]. An important property of PAN is associated with the absorption of toxic metal ions [4]. Polystyrene (PS) is also a synthetic thermoplastic polymer [5, 6]. It is transparent, hard, stiff, and one of the most widely used plastic. It is an economical polymer having scale of production of several million tons per year. Blending of polyacrylonitrile

and polystyrene is known in literature [7, 8]. PAN/PS blends have shown fine dispersion, thermal stability, and superior physical properties. PAN and carbon nanomaterial-based composite have been used for high performance industrial applications such as absorbent, electrode, gas storage media, catalyst support, etc. [9, 10]. Correspondingly, polystyrene reinforced with graphene and carbon nanotube possess high electrical conductivity, thermal stability, and mechanical strength to be employed in technical applications [11, 12]. Tribological behavior of polystyrene-acrylonitrile mixed with carbon nanotube has been reported [13, 14]. However, not much literature is available regarding the PAN/PS blends with filled carbonaceous nanofillers. Carbon nanoparticle (CNP) is a unique and less studied form of carbon nanomaterials [15]. CNP is a zero-dimensional nanocarbon having diameter <10 nm. It consists of multilayer fullerene-like structure [16]. CNP is capable of substituting other nanocarbon nanofillers such as carbon nanotube, graphene, graphite, owing to range of exceptional structural and physical properties. CNP has been reinforced in polystyrene, polyamide, polyaniline, polypyrrole, and polythiophene matrices [17]. CNP-based materials have found applications in batteries, supercapacitor, solar cell, and biomedical fields. To the best of knowledge, CNP nanofiller has not been reinforced in PAN or PAN blends. In this study, acid functional carbon nanoparticle has been reinforced in novel polyacrylonitrile/polystyrene blend to form polyacrylonitrile/polystyrene/carbon nanoparticle (PAN/PS/CNP) nanocomposite foams. PAN foam has been reported for various technical applications such as solar cell, fuel cell, Li-ion battery, sensor, biomedical, and removal of environmental toxicity [18-20]. However, PAN/PS/CNP nanocomposite foams is not reported before. This study reports on structure, morphology, compression mechanical properties, thermal stability, non-flammability, water uptake, and toxic metal ion removal of novel PAN/PS/CNP nanocomposite foams.

EXPERIMENTAL

Chemicals

Carbon nanoparticle (CNP, diameter 70-100 nm), polyacrylonitrile (PAN, average $M_w \sim 1.5 \times 10^5$), polystyrene (PS, $M_w 35 \times 10^3$), 2'-azobisisobutyronitrile (AIBN, 98%), and N,N-dimethylformamide (DMF, 99%) were supplied by Aldrich.

Instrumentation

Fourier transform infrared (FTIR) spectra were recorded using Excalibur Series FTIR Spectrometer, Model No. FTSW 300 MX manufactured by BIO-RAD. The compression strength and modulus were measured using universal electronic tensile machine (Shimadzu, Japan). The compressive rate was 2 mm/min according to ASTM D 1621-94. Morphology was investigated by FEI Nova 230 field emission scanning electron microscope (FE-SEM). The samples were cryogenically fractured in liquid nitrogen before analysis. Transmission electron microscopy (TEM) was performed with LEO 912 Omega instrument at 120 kV. The ultrathin sections were prepared using an Ultracut E ultramicrotome diamond knife. Thermal stability of materials was determined by NETZSCH thermogravimetric analyzer (TGA), model no. TG 209 F3, using 1-5 mg of the sample in Al_2O_3 crucible. The heating rate of 10 °C/min was maintained up to 800 °C under nitrogen flow rate of 30 mL/min. UL-94 test was executed ($100 \times 10 \times 5 \text{ mm}^3$) according to

ASTM D635-77. For moisture absorption, the test samples were placed in deionized water (24, 48, 72 & 96 h) after dry weighing. The specimens were removed from water and again weighed. The increase in mass was calculated for assessment of water absorption percentage. The change of Pb^{2+} ion concentration in aqueous solution was determined using atomic adsorption spectrophotometer (AAS, SHIMADZU AA-6701F). The time-dependent metal ion removal capacity (q_t) of Pb^{2+} ions was studied. The stock solution of Pb^{2+} ions was prepared by dissolving metal acetate salt in deionized water. Initial Pb^{2+} concentration of Pb^{2+} ion was 5 mmol metal ion/L (1000 mg/L) at pH 3. The pH was adjusted using 0.1M NaOH solution. The foam sample was placed in aqueous solution of 100 mL with desired initial Pb^{2+} ion concentration. All the experiments were carried out at room temperature.

Functionalization of carbon nanoparticle (CNP)

In 100 mL THF, 1g carbon nanoparticle was dispersed. Then, 50 mL conc. HNO_3 was added drop wise to the mixture with continuous stirring of 2h (70 °C). The product was washed with excess amount of distilled water to obtain the functional nanofiller. FTIR (KBr) spectrum reveal peaks at 3405 cm^{-1} (O–H stretching vibration), 1635 cm^{-1} (O–H bending vibration), 1721 cm^{-1} (C=O stretching vibration), 2890 cm^{-1} (C–H stretching vibration), and 1270 cm^{-1} (C–O) [21].

Preparation of polyacrylonitrile/polystyrene (PAN/PS) blend

0.9g polyacrylonitrile was dissolved in 20 mL DMF with continuous stirring of 24 h at 60 °C. 0.1g PS was also dissolved in 0.3 mL DMF separately. For blend formation, PAN and PS solutions were mixed and refluxed at 120 °C for 3h. The mixture was caste in glass Petri dishes at 80 °C for 24 h [22]. FTIR spectrum show peaks around 3010 (aromatic C–H stretching vibration), 2911 cm^{-1} (aliphatic C–H stretching vibration), 2231 cm^{-1} (–C≡N stretching vibration), 1450 cm^{-1} (C–N stretching vibration), and 1551, 1647 cm^{-1} (C=C stretching vibration).

Preparation of polyacrylonitrile/polystyrene/carbon nanoparticle (PAN/PS/CNP) nanocomposite

The PAN/PS blend solution was prepared according to the procedure mentioned in above Section 2.4. After refluxing the blend mixture at 120 °C for 3 h, CNP nanofiller (0.1-3 wt.%) was added to the mixture. The reaction mixture was futher refluxed for 2h at 120 °C. The films were caste in Petri dishes at 80 °C for 24h under vacuum (Fig. 1).

Preparation of PAN/PS/CNP Foam

Initially, PAN/PS/CNP nanocomposite (0.1-3 wt.%) was prepared as mentioned in above Section 2.5. Then solution of 3 wt.% foaming agent AIBN was prepared in 1mL DMF and added to the nanocomposite mixture. The mixture was refluxed at 120 °C for 1h. The reaction mixture was poured in Petri dish and allowed to dry at 80 °C for 12 h. The dried nanocomposite films were finally hot pressed. During pressing, foaming agent was decomposed to form foam structures [23]. FTIR peaks identified as 3231 cm^{-1} (O–H stretch), 3001 cm^{-1} (aromatic C–H stretch), 2906 cm^{-1} (aliphatic C–H stretching vibration), 2229 cm^{-1} (–C≡N stretching vibration), 1441 cm^{-1} (C–N stretching vibration), and 1530, 1610 cm^{-1} (C=C stretching vibration).

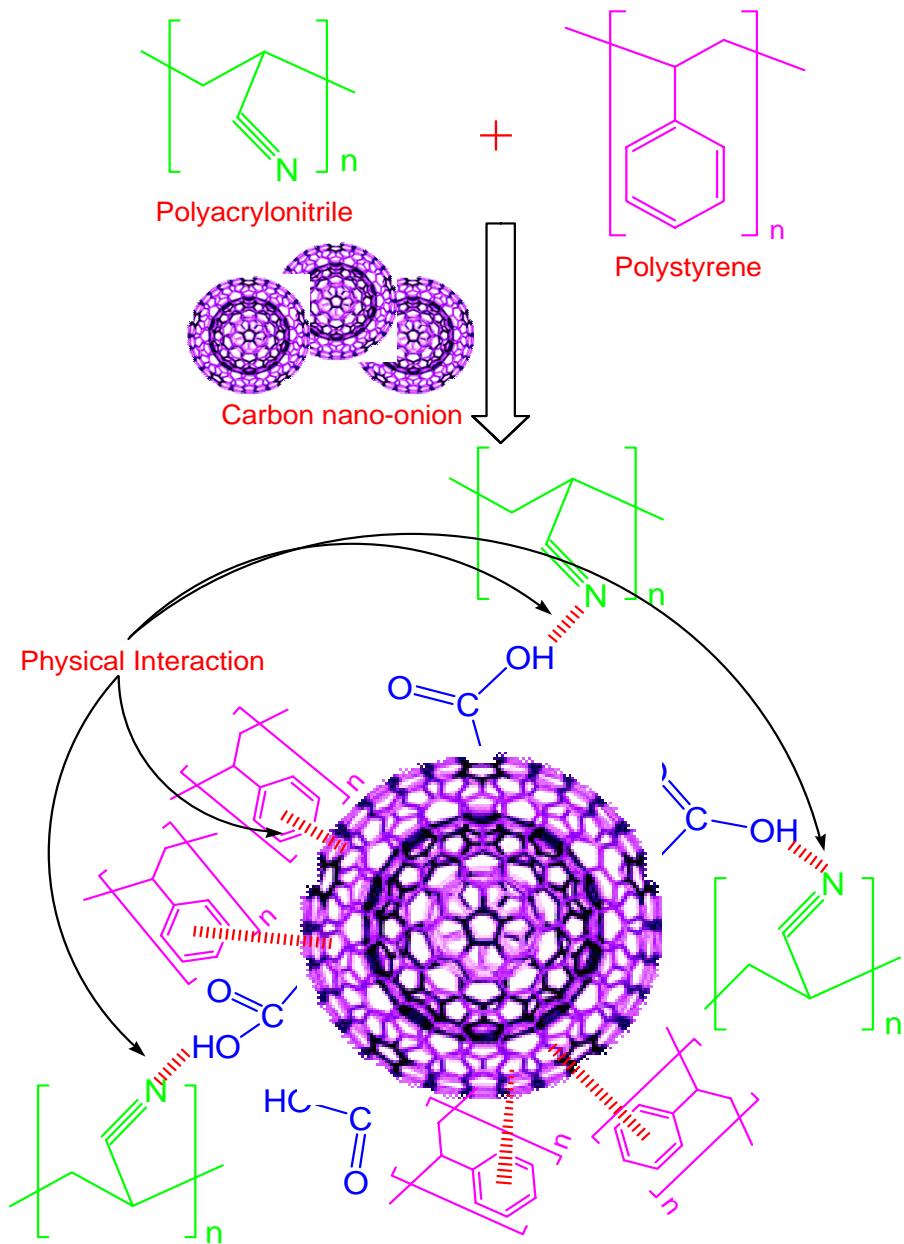


Fig. 1. Schematic for the formation of PAN/PS and CNP nanocomposite

RESULTS AND DISCUSSION

FTIR analysis

Fig. 2 reveals FTIR spectral features of PAN/PS and PAN/PS/CNP 3 Foam. The nitrile group of PAN backbone appeared at 1450 cm^{-1} . The $-C≡N$ peak at 2231 cm^{-1} also show the presence of PAN structure (Fig. 2A). The PAN/PS also depicts aromatic C-H stretching vibration at 3010 cm^{-1} due to polystyrene rings in the spectrum. The peaks at 1551 and 1647 cm^{-1} were

attributed to aromatic C=C stretching vibrations. The aliphatic protons also appeared at 2911 cm⁻¹. Fig. 2B shows the FTIR spectrum of nanocomposite foam. After the incorporation of functional CNP nanofiller, the hydroxyl O-H group was observed at 3231 cm⁻¹. Moreover, -C≡N group of PAN was found at lower wavenumber 2229 cm⁻¹ due to hydrogen bonding interaction with the hydroxyl group of functional CNP. The C-N stretching vibration was also found at lower position (1441 cm⁻¹). Due to ring stacking interaction, aromatic protons (3001 cm⁻¹) and C=C (1530, 1610 cm⁻¹) stretching vibrations were also shifted to lower wavenumber. However, separate carbonyl peak (1700 cm⁻¹) cannot be observed in the spectrum may be due to interaction between the matrix and nanofiller and possible overlapping with aromatic C=C stretching vibration. Consequently, spectral data offered sufficient evidences regarding the incorporation of carbon nanoparticle in polymer framework and also interaction between the matrix-nanofiller.

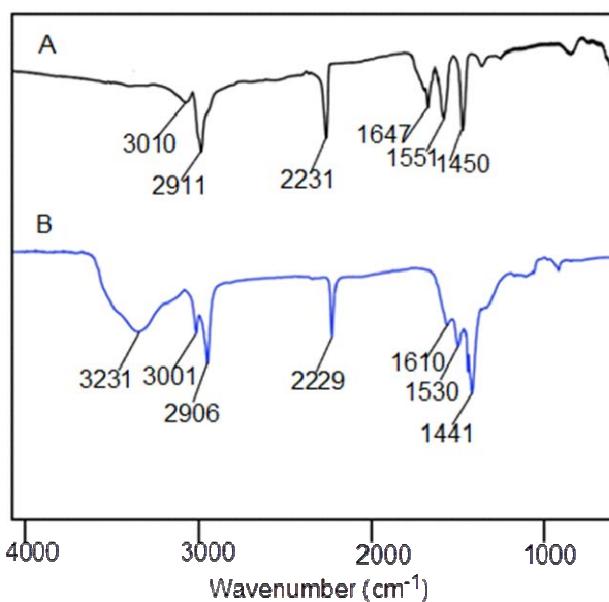


Fig. 2. FTIR spectra of (A) PAN/PS and (B) PAN/PS/CNP 3 Foam

Microstructure study

Field emission scanning electron microscopy (FESEM) was used to explore the microstructure of polyacrylonitrile/polystyrene blend and PAN/PS/CNP nanocomposite. The morphology of PAN/PS was found spongy in nature with small pores (Fig. 3A). Unsurprisingly, neat polymer blend foam had ill-defined cell structure. Addition of 0.1 wt.% nanofiller not caused any substantial change in the morphology of PAN/PS blend (Fig. 3B). However, 0.3 wt.% nanofiller addition instigated visible change in morphology relative to the neat blend (Fig. 3C). Morphology of PAN/PS/CNP 1 Foam seems to have spongy, branched, and porous morphology (Fig. 3D). Additional inclusion of nanofiller up to 3 wt.% also depicts same branched spongy morphology (Fig. 3E).

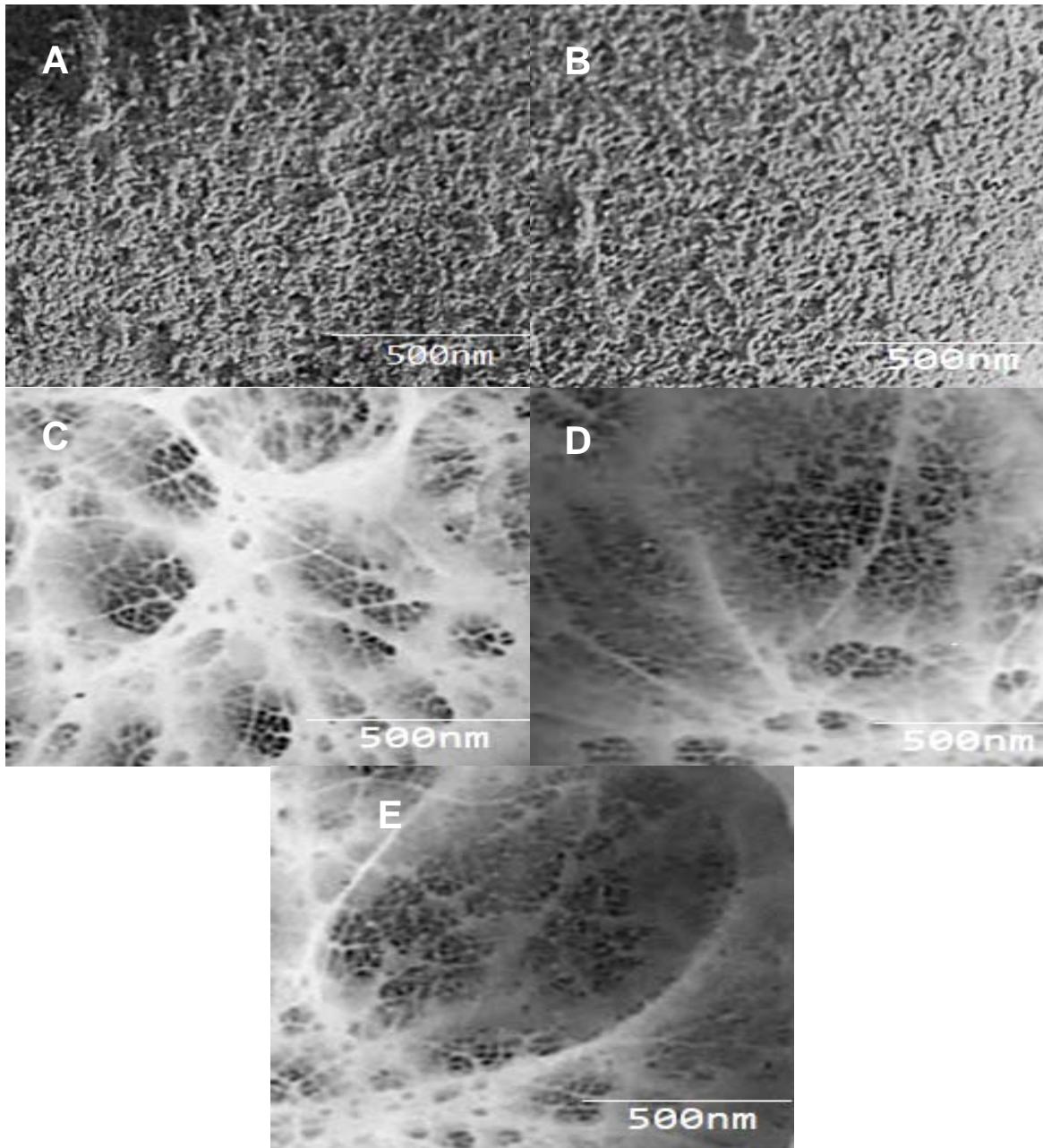


Fig. 3. FESEM micrographs of (A) PAN/PS Foam; (B) PAN/PS/CNP 0.1 Foam; (C) PAN/PS/CNP 0.3 Foam; (D) PAN/PS/CNP 1 Foam; and (E) PAN/PS/CNP 3 Foam

The unique morphology was attributed to physical interaction between the matrix and nanofiller and synergetic effect in the blended polymers and functional CNP leading to the formation of homogenously aligned framework [24]. Nevertheless, the cell structure of foams was not evident in the FESEM micrographs. TEM micrographs of PAN/PS Foam (Fig. 4A), PAN/PS/CNP 0.3 Foam (Fig. 4B), and PAN/PS/CNP 3 Foam (Fig. 4C) show exclusive nanocellular morphology. The cells were square shaped and homogeneously dispersed in the foam matrix. Perfectly aligned and uniform cell structure was observed in TEM micrographs.

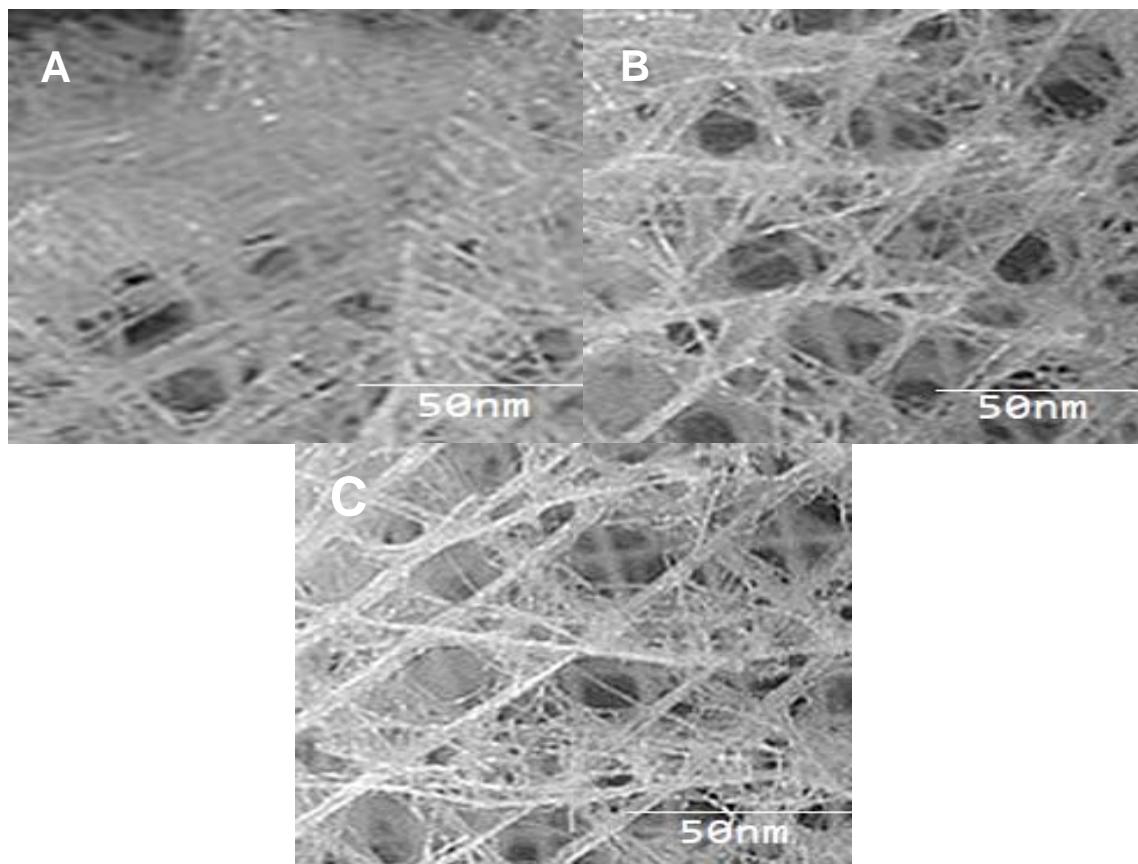


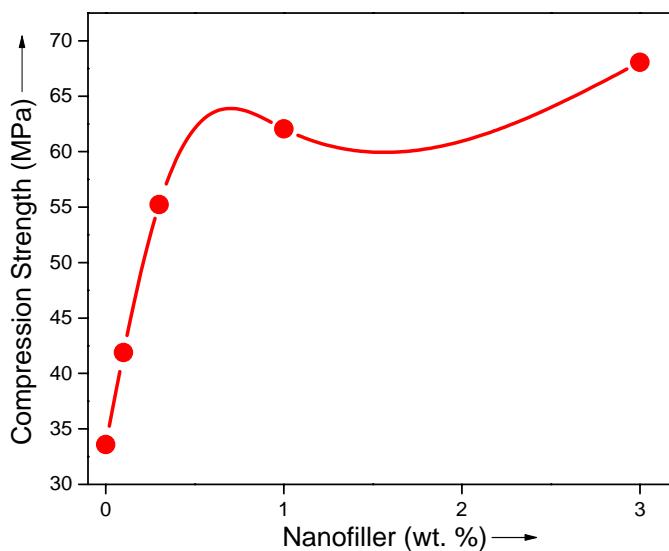
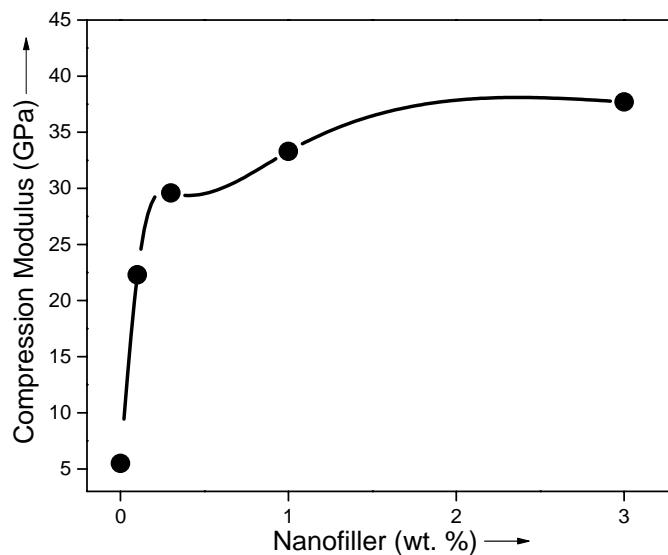
Fig. 4. TEM micrographs of (A) PAN/PS Foam; (B) PAN/PS/CNP 0.3 Foam; and (C) PAN/PS/CNP 3 Foam

Compression strength and modulus of PAN/PS Foam and PAN/PS/CNP Foam

Table 1 shows compression strength and compression modulus of pristine PAN/PS blend and PAN/PS/CNP nanocomposite foam. Fig. 5 and Fig. 6 show increase in compression mechanical properties with CNP addition. Neat PAN/PS sample possess compression strength and modulus of 33.5 MPa and 5.5 GPa, respectively. PAN/PS/CNP 0.1 Foam with 0.1 wt.% nanofiller yield compression strength of 41.8 MPa. Inclusion of 0.3 wt.% nanofiller enhanced the compression strength upto 55.2 MPa. Further addition of 0.6 and 1 wt.% nanofiller enhanced the strength to 57.7 and 62.1 MPa. PAN/PS/CNP 3 Foam showed highest compression strength of 68.2MPa. In this way, there was 51% increase in the compression strength of 3 wt.% foam compared with the neat polymer foam. The effect of nanocarbon nanofillers has already been observed for enhanced strength properties [25, 26]. Similar increasing trend was observed for the compression modulus of nanocomposite foams. Compression modulus of 0.1 wt.% loaded foam was 22.3 GPa. Reinforcement from 0.3-3 wt.% increased the modulus values from 29.6 to 37.7 GPa. There was 85% increment in the compression modulus of PAN/PS/CNP 0.3 Foam relative to the neat blend foam. Increase in mechanical properties of nanocomposite foams with nanocarbon addition follows the literature trend observed [27, 28]. Carbon nanoparticles were found advantageous to enhance the mechanical properties of novel foams relative to graphene and carbon nanotube-coated polymer foams.

Table 1. Mechanical properties and density of PAN/PS Foam and PAN/PS/CNP Foam

Sample	Compression Strength (MPa)	Compression Modulus (GPa)	Density (mgcm^{-3})
PAN/PS	33.5	5.5	0.5
PAN/PS/CNP 0.1 Foam	41.8	22.3	0.9
PAN/PS/CNP 0.3 Foam	55.2	29.6	1.1
PAN/PS/CNP 0.6 Foam	57.7	30.1	1.4
PAN/PS/CNP 1 Foam	62.1	33.3	1.7
PAN/PS/CNP 3 Foam	68.2	37.7	1.9

**Fig. 5.** Compression strength of PAN/PS Foam and PAN/PS/CNP Foam**Fig. 6.** Compression modulus of PAN/PS Foam and PAN/PS/CNP Foam

Thermogravimetric and non-flammability measurements

Table 2 shows thermogravimetric analysis of blend and PAN/PS/CNP nanocomposite foams. The thermograms are presented in Fig. 7. The PAN/PS blend had initial weight loss (T_0), 10% degradation temperature (T_{10}), and maximum decomposition temperature (T_{max}) of 411, 418, and 459 °C, respectively. Thermal properties of neat blend were found superior relative to neat PS or PAN owing to blending effect of the two polymers. Incorporation of 0.1 wt.% nanofiller enhanced the T_0 , T_{10} , and T_{max} to 501, 523, and 555 °C, respectively. The 0.3wt.% reinforcement further enhanced the T_0 , T_{10} , and T_{max} to 511, 530, and 559 °C. The thermal properties were raised constantly with CNP loading and attained T_0 of 529 °C, T_{10} of 544 °C and T_{max} of 568 °C for PAN/PS/CNP 1 Foam. Greatest improvement in thermal properties were achieved with 3wt.% nanofiller loading ($T_0 = 541$, $T_{10} = 559$, and $T_{max} = 574$). Consequently, initial and maximum weight loss values were improved by 24% and 20% respectively, relative to neat blend foam. It seems that the nanofiller loading developed physical interaction with the foam matrix and led to unique microcellular morphology. The stable matrix-nanofiller framework in turn improved the thermal stability and mechanical strength of PAN/PS/CNP foams. Table 2 also illustrates non-flammability data of the samples. According to UL-94 test, V-0 rating was obtained for all the nanocomposite foams. Carbon nanoparticle addition was found to constantly increase the char yield of the nanocomposite. According to literature, high char yield of nanocomposite foams (45-57%) points to non-flammability capability of the samples [29-31]. Inclusion of 3 wt.% of nanofiller considerably enhanced the char yield of neat PAN/PS blend (22%) up to 57%. The enhanced char yield can be attributed to inherent flame retardancy of carbon nanoparticle and formation of a stable matrix-nanofiller foam network resilient to heat. Thus, the nanocomposite structure and non-flammability behavior of CNP nanofiller supported the flame retardancy of PAN/PS/CNP foams [23, 32]. Compared with the literature polyamide and graphene systems, novel PAN and carbon nanoparticles have shown comparable thermal and non-flammability properties [31]. Thus, the effect of carbon nanoparticle has been identified in improving the heat constancy of the materials.

Table 2. Thermal analyses and flammability data of PAN/PS Foam and PAN/PS/CNP Foam

Sample	T_0 (°C)	T_{10} (°C)	T_{max} (°C)	Char yield (%)	UL-94
PAN	265	276	399	44	-
PS	319	370	390	10	-
PAN/PS	411	418	459	22	V-0
PAN/PS/CNP 0.1 Foam	501	523	555	45	V-0
PAN/PS/CNP 0.3 Foam	511	530	559	47	V-0
PAN/PS/CNP 0.6 Foam	517	539	564	49	V-0
PAN/PS/CNP 1 Foam	529	544	568	50	V-0
PAN/PS/CNP 3 Foam	541	559	574	57	V-0

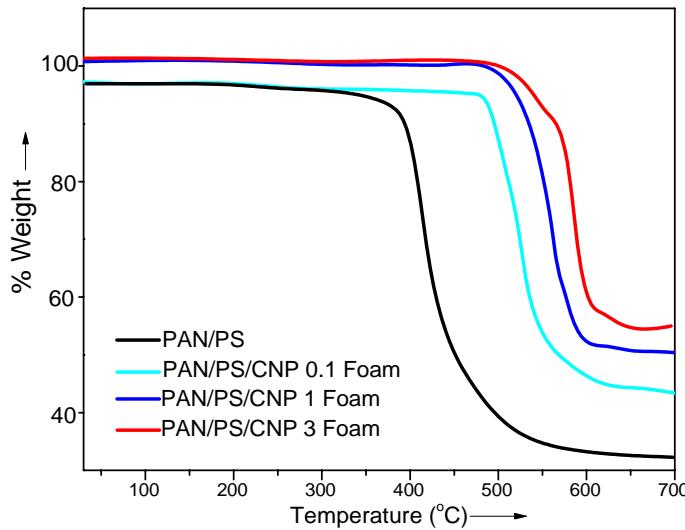


Fig. 7. TGA curves of PAN/PS and PAN/PS/CNP nanocomposite foam at 10 °C/min (N2)

Moisture absorption analysis

Table 3 shows moisture absorption test performed using a simple method. First, the initial dry weight of samples was measured. Then, the samples were immersed in de-ionized water for the desired period of time (24, 48, 72, 96 h). The study was carried out at room temperature (27 °C). After the completion of test period, foam specimens were taken out from water and dried. The dried samples were weighed to examine the weight difference and % water absorption (Eq. 1).

$$\text{Increase in weight (\%)} = (W-D)/D \times 100 \quad (1)$$

Where W = wet weight and D = dry weight. After 24 h, neat foam, PAN/PS/CNP 0.1 Foam, PAN/PS/CNP 0.3 Foam, PAN/PS/CNP 1 Foam, and PAN/PS/CNP 3 Foam showed water absorption of 7.31, 9.31, 12.1, 18.3, and 20.6%, respectively. Increase in soaking time period considerably enhanced the moisture absorption by foams. The overall trend obtained is shown in Fig. 8. Maximum moisture absorption was observed after 96 h as 53.8, 62.3, 69.9, 72.3, and 77.4% respectively for neat foam, PAN/PS/CNP 0.1 Foam, PAN/PS/CNP 0.3 Foam, PAN/PS/CNP 1 Foam, and PAN/PS/CNP 3 Foam. Thus, higher water absorption was observed with 3 wt.% nanofiller loading (96 h) in the foam sample i.e. 30% higher than the pristine foam. Increase in water absorption with nanofiller loading can be attributed to the presence of hydrophilic nature of acid functional CNP in the foam structure [33-35]. The Young's contact angle of neat blend and nanocomposite was also measured to perceive the water absorption trend. The contact angle of the blend was low (80°). Varying the weight fraction of nanofiller blended with polymers systematically increased the contact angle to about 120° . The increase in contact angle with nanofiller loading also points to hydrophilic nature of functional CNP.

Table 3. %Water absorption and Pb²⁺ ion removal efficiency of PAN/PS Foam and PAN/PS/CNP Foam

Time (h)	% Water absorption				
	PAN/PS Foam	PAN/PS/CNP 0.1 Foam	PAN/PS/CNP 0.3 Foam	PAN/PS/CNP 1 Foam	PAN/PS/CNP 3 Foam
24	7.31	9.31	12.1	18.3	20.6
48	17.4	20.8	26.1	29.3	39.1
72	38.1	45.3	49.1	50.7	59.9
96	53.8	62.3	69.9	72.3	77.4
Metal ion removal capacity (mg/g)	90	141	149	153	166
Removal efficiency (%)	45	77	90	92	99
Contact angle (°)	80	83	98	110	120

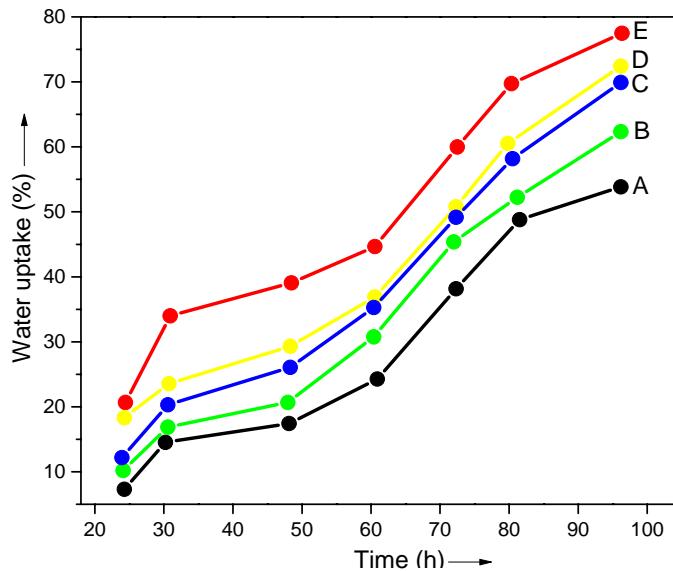


Fig. 8. Water uptake characteristics of (A) PAN/PS Foam; (B) PAN/PS/CNP 0.1 Foam; (C) PAN/PS/CNP 0.3 Foam; (D) PAN/PS/CNP 1 Foam; and (E) PAN/PS/CNP 3 Foam

Ion removal

Time-dependent amount (q_t) of Pb²⁺ ions removed by PAN/PS/CNP 0.3 Foam and PAN/PS/CNP 3 Foam from aqueous solution is shown in Fig. 9. The Eq. 2 was used to calculate q_t (mg/g) for Pb²⁺ ions:

$$q_t = \frac{(C_0 - C)V}{B} \quad (2)$$

Where C_0 = initial Pb²⁺ ion concentration (mg/l); C = residual Pb²⁺ ion concentration (mg/l); V = volume of solution (l); and B = weight of nanocomposite foam (g). The initial Pb²⁺ ion

concentration was 100 mg/L at pH 3. The removal efficiency of foams was also calculated using Eq. 3:

$$\text{Removal efficiency (\%)} = \frac{C_0 - C}{C_0} \quad (3)$$

Table 3 shows Pb²⁺ ions removal efficiency. It was found that the nanofiller content directly influenced the q_t and removal efficiency of Pb²⁺ ions from aqueous solution. The amount of Pb²⁺ ions removed by 3 wt.% nanofiller foam was much higher relative to 0.3 wt.% foam sample. Higher nanofiller loading offered better adsorption sites for Pb²⁺ ions. Up to 30h, there was continuous increase in q_t of Pb²⁺ ions due to homogenous adsorption and distribution of adsorption sites. From 40-50h, there was slight decrease in the ion adsorption owing to hindrance of active sites with the adsorbed ions. As the time passes, the ion adsorption was again increased owing to the accumulation of active sites for ion removal. Consequently, prolonged exposure of foams to metal ions enhanced the removal capacity. Consequently, PAN/PS/CNP 3 Foam revealed removal capacity of 166 mg/g and removal efficiency of 99%. The results depict positive impact of CNP on the removal of toxic metal ion.

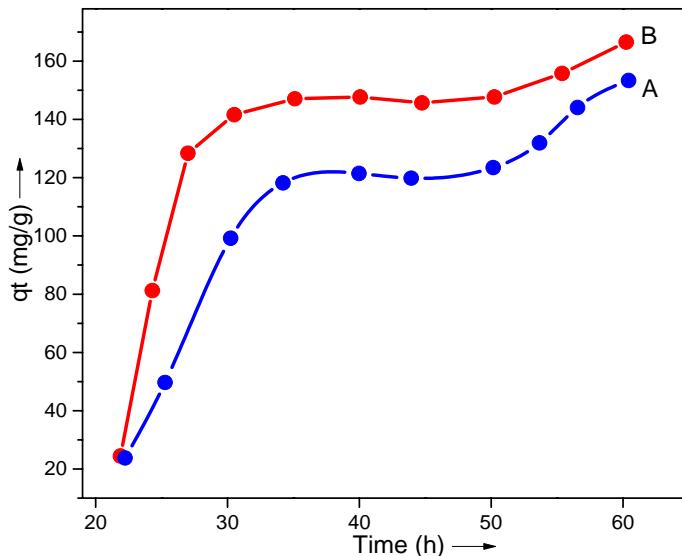


Fig. 9. Time-dependent amount (q_t) of Pb²⁺ ions removed by (A) PAN/PS/CNP 0.3 Foam and (B) PAN/PS/CNP 3 Foam in aqueous solutions with initial Pb²⁺ ion concentration of 100 mg/L at pH 3

Testing of experimental results were also performed by adsorption isotherm method. According to Redlich-Petersen isotherm (Eq. 2):

$$f = Kce M / (1 + (Kce)^\beta) \quad (2)$$

f = ionic species sorbed per gram of material (mmol g⁻¹); ce = equilibrium concentration of ions (mmol L⁻¹); M = maximum adsorption (mmol g⁻¹); β = degree of heterogeneity; K=constant. Fig. 10 graphically shows isotherm and curve of adsorption model approached. The isotherm has shown changes in surface sorption properties during ion binding. The Redlich-Petersen curve gave much better fit at higher concentrations of Pb. It was found that the ionic species bind with the most favorable sites for multi-layer adsorption [36]. The adsorption isotherm of PAN/PS/CNP 3 Foam

revealed better removal of species owing to appropriate adsorption sites on the nanocomposite surface. The effect of metal ion removal was found similar to the removal efficiency of the reported nanocarbon nanocomposite [37-39].

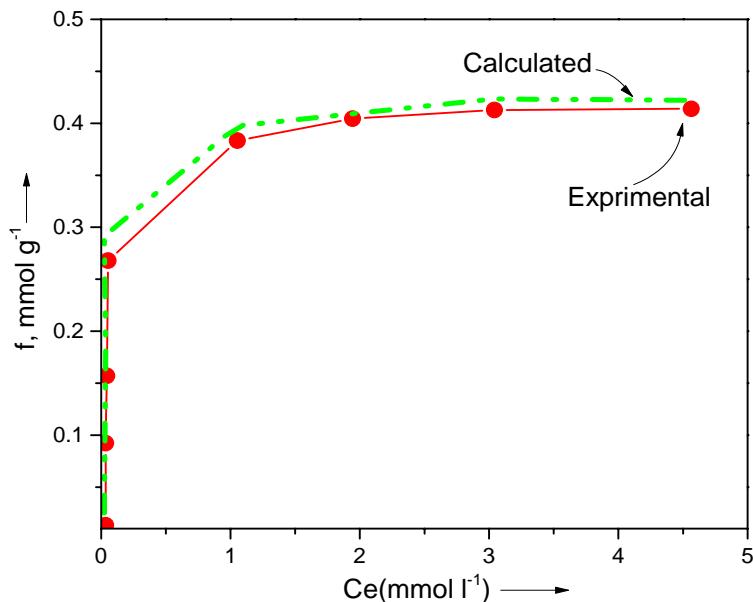


Fig. 10. Adsorption isotherms of Pb^{2+} ions for PAN/PS/CNP 3 Foam

CONCLUSIONS

Novel polyacrylonitrile/polystyrene/carbon nanoparticle-based foams were designed, fabricated, and investigated using FTIR, FESEM, TEM, TGA, UL94, compression tests, water absorption, and toxic ion removal. FTIR and microscopic studies have shown that the carbon nanoparticle nanoparticles were successfully incorporated and dispersed in the blend matrix through physical interaction. Among nano-cellular foams, 3 wt.% CNP loading had unique nanocellular morphological pattern, high foam density, and compressive mechanical properties. High heat stability and flame resistance were also observed via superior thermal properties, char yield, and V-0 rating. In addition, the nanocomposite foam with high CNP content showed fine moisture uptake and Pb^{2+} ions removal capacity/efficiency. Thus, the cautiously designed blend structure and functional nanofiller led to the development of high performance strengthened and heat and flame retardant nanocellular foams for metal ion removal.

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