

The Ti-MWW catalyst – its characteristic and catalytic properties in the epoxidation of allyl alcohol by hydrogen peroxide

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Ti-MWW, one of the latest titanium-silicalite catalysts, has been prepared by direct hydrothermal synthesis using hexamethyleneimine as a structure-directing agent. The characteristic of the catalyst was performed by means of the following methods: XRD, SEM, IR, UV-vis and X'Ray microanalysis. The catalytic properties of Ti-MWW have been compared with those of the conventional titanium-silicalites TS-1 and TS-2 in the epoxidation of allyl alcohol with hydrogen peroxide. The process has been described by the following main functions: the selectivity of the transformation to glycidol in relation to allyl alcohol consumed, the conversions of the substrates (allyl alcohol and hydrogen peroxide) and the selectivity of the transformation to organic compounds in relation to hydrogen peroxide consumed.

Keywords: Ti-MWW, titanium-silicalites, glycidol, liquid phase epoxidation, hydrogen peroxide

INTRODUCTION

An important structural parameter of the zeolites is the size of the pore opening through which organic molecules diffuse into channels and cages of zeolite to active centers. It limits the possibility of using the catalysts in organic reactions, especially in epoxidation processes¹.

Since titanium-containing zeolite with MFI structure (TS-1) was synthesized using hydrothermal synthesis method by Taramasso et al.² in 1983, much attention has been focused on the TS-1/H₂O₂ reaction system owing to its environmentally benign and outstanding oxidation selectivity. The next synthesized titanium-silicalite catalysts were: TS-2 (the pore size 0.5 nm – the same as TS-1), Ti-Beta (the pore size 0.7 nm). Silicates Ti-MCM-41 and Ti-MCM-48 (the pore size depends on template, the catalysts synthesized by us had the pore size 1.3 nm) have similar properties. These catalysts were discovered because the researchers were looking for the catalyst having larger pore size than TS-1 and they will be used for the epoxidation of bulky and branched organic molecules. Ti-MWW³ belongs to the latest titanium-silicalite catalysts.

The new titanium-silicalite Ti-MWW catalyst structure is built from a lamellar precursor undergoing

dehydroxylation upon calcinations between the layered sheets (Figure 1). Besides two-dimensional sinusoidal channels of 10-membered rings (MR) running throughout the structure parallel to the *ab*-plane, the MWW structure contains an independent channel system which is comprised of large supercages 0.7 × 0.7 × 1.8 nm. The supercages turn out to be pocket or cup moieties 0.7 × 0.7 nm at the crystal exterior. This may provide potential opportunities for a wide variety of applications in petrochemical and fine chemical industries^{4–5}.

Ti-MWW is the active catalyst in: epoxidation of linear alkenes⁶ and allyl alcohol⁷ with hydrogen peroxide, oxidation of *cis*- and *trans*-2-hexenes^{8,9}, epoxidation of bis(allyl) ether (DAE) to allyl glycidyl ether (AGE) and diglycidyl ether (DGE) with aqueous hydrogen peroxide¹⁰ and liquid-phase ammoxidation of cyclohexanone to oxime¹¹.

The aim of this work was the synthesis of Ti-MWW. The next step was the application of Ti-MWW catalyst in the epoxidation of allyl alcohol over Ti-MWW catalyst and comparison of these results with those obtained over TS-1 and TS-2 catalysts.

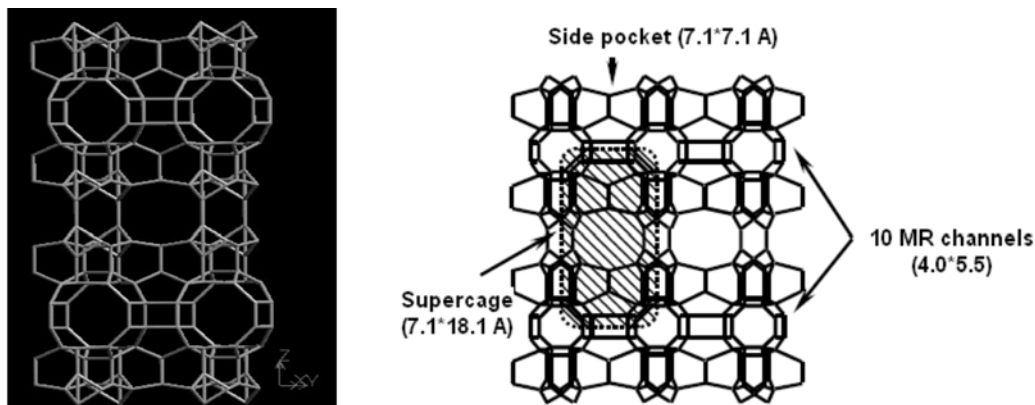


Figure 1. Topology of Ti-MWW zeolite⁵

EXPERIMENTAL

Preparation of Ti-MWW catalyst and its characteristics

For the synthesis of titanium-silicalite Ti-MWW catalyst the following raw materials: tetrabutyl o-titanate (TBOT, $\text{Ti}(\text{OBu})_4$, 95%, Fluka), hexamethylenimine (HMI, $\text{C}_6\text{H}_{13}\text{N}$, >98%, Merck), boric acid (H_3BO_3 , 99.5%, Chempur), and fumed silica (Cab-o-sil M5, Biesterfeld)) were applied.

The Ti-MWW catalyst was synthesized by the method of Wu et al.¹²: hexametyleneimine (HMI) was dissolved in deionized water at room temperature. This solution was divided into two equal parts, to one tetrabutyl o-titanate $\text{Ti}(\text{OBu})_4$ was added, and to the other boric acid was added under vigorous stirring. The stirring was maintained for 30 min to hydrolyze tetrabutyl o-titanate. Fumed silica (Cab-o-sil M5) was divided into two equal parts which were added gradually to the solutions containing Ti (titanium) and B (boron), respectively. A further stirring for 1 hour was allowed to form two homogeneous gels. The gels were then mixed together and stirred for 1.5 hours to obtain a gel. The resulting gel was transformed into a PTFE-lined autoclave (Berghof) and heated with a stirring rate of 150 rpm at 170°C (443K) for 14 days. After cooling, the solid product was filtered off and washed with deionized water to pH < 9. The product was then dried at 50°C (323K) for 2 – 3 days. To remove boron and extraframework Ti species, an acid treatment with 2M HNO_3 solution was carried out at 100°C (373K) for 20 hours at a solid-liquid ratio of 1g to 20 ml. Then the product was dried at 50°C (323K) for 2 – 3 days and calcined at 530 – 550°C to burn off any residual organic species.

The stages of the synthesis of Ti-MWW catalyst can be generally presented as follows:

the synthesis of Ti-MWW gel → crystallization of the gel → washing with deionized water to pH < 9 and drying → acid treatment to remove boron → calcination.

The Ti-MWW catalyst was investigated by means of the XRD method (the X-ray diffraction spectroscopy) on X'Pert PRO Philips diffractometer using $\text{CuK}\alpha$ radiation. This method allows to establish the type of crystalline structure of the catalyst. The XRD patterns of the Ti-MWW catalyst (Figure 2) were totally consistent with the XRD patterns of Ti-MWW presented in literature⁴. The peaks at $2\theta = 5 - 7^\circ$ are characteristic of a lamellar structure along *c*-direction – Figure 2a. Other peaks are related with the crystalline sheets parallel to the *ab*-planes. Peak 6.57° disappeared after acid treatment and calcination, while the peaks of 7.27° and 8.02° remained unchanged. This indicates that calcination only affected the *c*-axis structure, that is, calcination caused the loss of the laminar structure and the formation of the ordered linkage between the sheets through oxygen bridges ($-\text{O}-$)⁴. The peaks $2\theta = 7 - 10^\circ$ and $2\theta = 25 - 27^\circ$ for the Ti-MWW after acid treatment and calcination are weaker than those of Ti-MWW after crystallization (Figure 2b). The lower crystallinity may arise from a small amount of framework defect sites generated by removing the organic species of B and Ti by the acid treatment⁴.

Scanning electron micrographs (SEM method) of the Ti-MWW catalyst are shown in Figure 3. The micro-

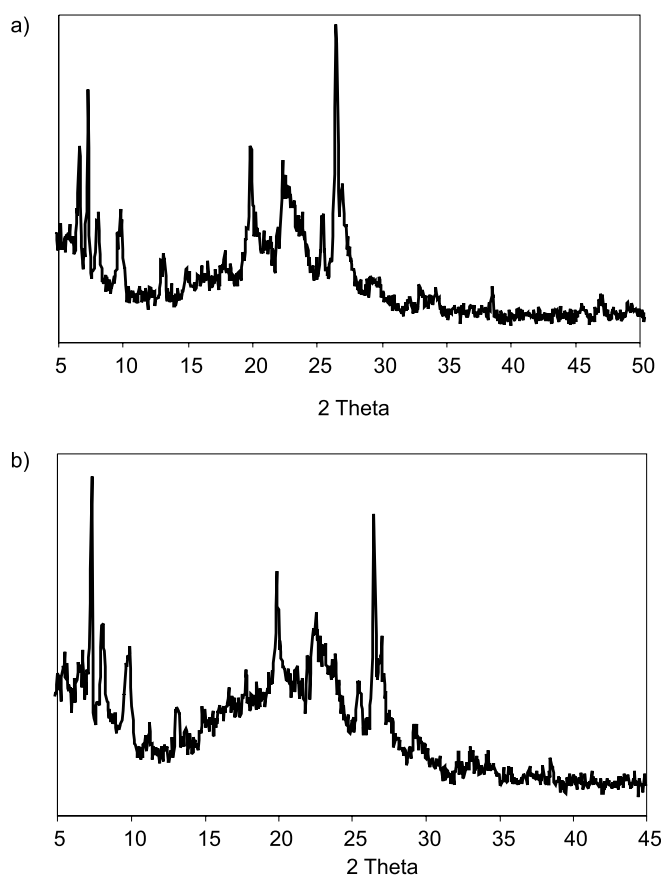


Figure 2. XRD patterns of the Ti-MWW obtained using hexamethylenimine as a template: a) after crystallization and b) after acid treatment and calcination

graphs were made on JOEL JSM-6100 instrument. All the crystals appeared to be thin plates that mainly exhibited the hexagonal morphology. The thin plates of the Ti-MWW samples are approximately $2.5\ \mu\text{m}$ in length and $0.3\ \mu\text{m}$ in thickness and form aggregates. The calcination does not change the shape and size of the crystals.

The IR investigations were made on Shimadzu FTIR-8100 spectrometer using the KBr pellet technique and are the same as in the literature⁴. The IR method is frequently employed to characterize the framework tetrahedral Ti in Ti-containing molecular sieves due to a characteristic band at $960\ \text{cm}^{-1}$, which is assigned to the vibration of Si-O-Ti. The Ti-MWW sample, especially after acid treatment, contained a high concentration of defect sites such as hydroxyl nests generated on the vacancy of the framework B and Ti (band around $3450\ \text{cm}^{-1}$) – Figure 4. The tetrahedral framework B develops two strong Si-O-B stretching bands: $1400\ \text{cm}^{-1}$ and $940\ \text{cm}^{-1}$, the latter of which overlaps the $960\ \text{cm}^{-1}$ band to hinder the IR characterization of framework Ti species but after acid treatment it disappears. In the region of the zeolite framework vibration, the bands showing a cut off at $1000 - 1300\ \text{cm}^{-1}$ is due to the Si-O-Si stretching vibration⁴.

The investigations by UV-VIS method was made on SPECORD M40 type V-530. As shown in Fig. 5, the UV-vis spectrum of Ti-MWW after acid treatment and calcination shows the main band at 260 nm. The band confirms the presence of Ti in tetrahedral coordination in the material. The other additional bands were not observed, especially the band around 330 nm, which confirms the presence of anatase in the sample.

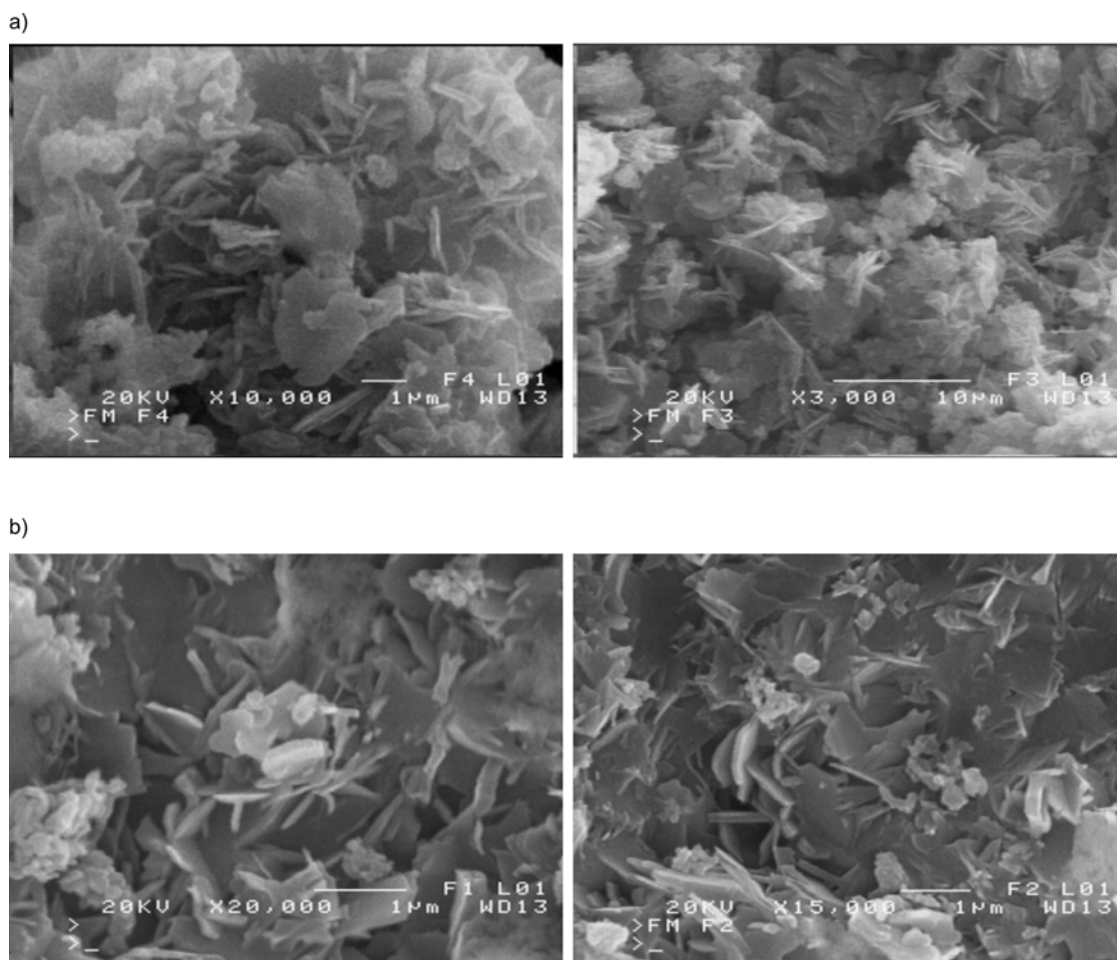


Figure 3. Scanning electron micrographs of the Ti-MWW catalyst obtained using hexamethyleneimine as a template: a) – after crystallization and b) – after acid treatment and calcination

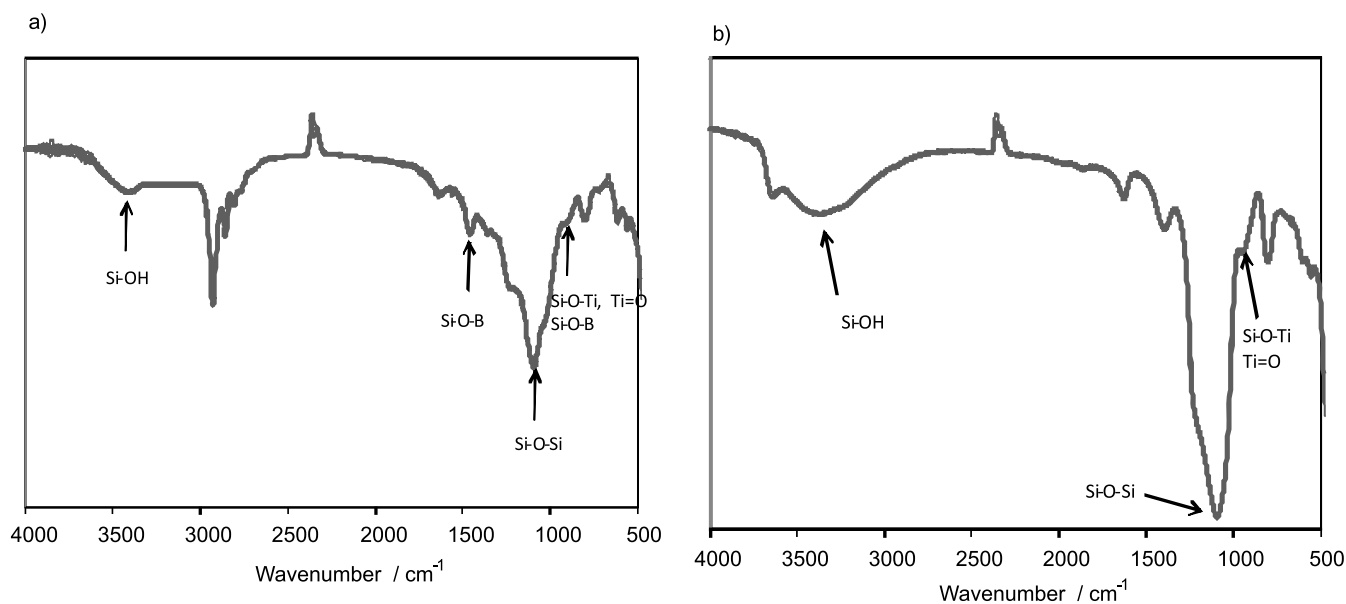


Figure 4. IR spectra of Ti-MWW catalyst obtained using hexamethyleneimine as a template: a) – after crystallization and b) – after acid treatment and calcination

The investigations by X-ray microanalysis was made on the Oxford X-ray analyzer ISIS 300. The amount of Ti in the sample after crystallization was 5.68 wt%, and after acid treatment and calcination 3.22 wt%.

Preparation of TS-1 and TS-2 catalysts

The TS-1 catalyst (to 2,50 wt% Ti) was prepared according to the method of Thangaraj et al.¹³ The detailed

characteristic of the catalyst was presented in our previous works^{14–15}.

The TS-2 catalyst (to 3,50 wt% Ti) was synthesized according to the method described by Reddy et al.¹⁶. The detailed characteristic of the catalyst was presented in our previous works^{15, 17}.

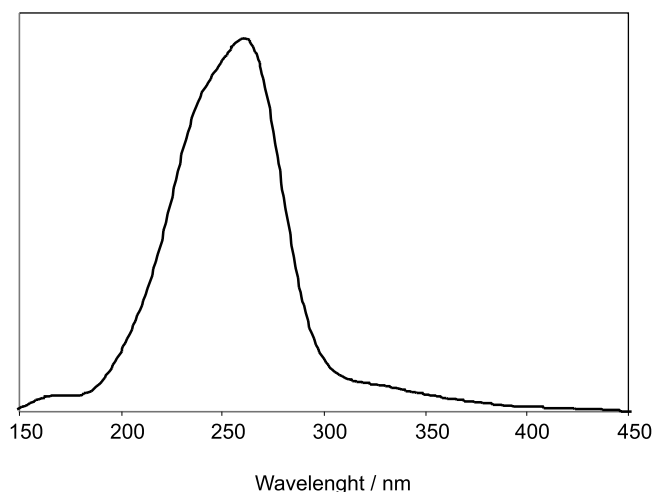


Figure 5. UV-vis spectra of the Ti-MWW catalyst after acid treatment and calcination

Epoxidation procedure, apparatus and analytical methods

In the epoxidation of allyl alcohol the following raw materials were used: allyl alcohol (98 wt%, Fluka), the TS-1, TS-2 and Ti-MWW titanium silicalite catalysts (prepared at the Institute of Organic Chemical Technology, West Pomeranian University of Technology, Szczecin), hydrogen peroxide (30 wt% water solution, P.O.Ch. Gliwice, Poland) and methanol (analytical grade., P.O.Ch. Gliwice, Poland).

The process of allyl alcohol epoxidation was carried out under atmospheric pressure in a glass reactor equipped with a reflux condenser, thermometer, mechanical stirrer and a dropping funnel. The substrates in proper amounts were introduced into the reactor in the following sequence: titanium-silicalite catalyst, allyl alcohol and methanol as a solvent. When the temperature reached a target value 30 wt% water solution of hydrogen peroxide was dropped in. The process was conducted for a certain amount of time, then the content of the reactor was weighed and the product was analysed. The technological parameters used in the epoxidation over TS-1, TS-2 and Ti-MWW catalysts were as follows: temperature 20°C, the molar ratio AA/H₂O₂ 1:1, methanol concentration 80 wt% and reaction time 3h.

In order to make mass balances of the syntheses performed, the following analyses were made: unreacted hydrogen peroxide was iodometrically determined¹⁸, glycerine formed in the process was potentiometrically determined¹⁹, the other products and the unreacted allyl alco-

hol were determined by gas chromatography. The chromatographic analyses were performed on a FOCUS apparatus equipped with a flame-ionization detector (FID), using a capillary column Quadrex 30m×250μm×0,25μm packed with methylsiloxane modified with phenyl groups. The parameters of chromatographic separation were as follows: helium pressure 50 kPa, sensitivity 10, the temperature of the sample chamber 150°C, detector temperature 250°C. The thermostat temperature was programmed in the following way: isothermally 40°C for 3 min, followed by an increase at the rate 10°C/min to 250°C, isothermally 250°C for 5 min, cooling to 60°C. After the calculation of the mass balance for each synthesis, the main functions describing the process were determined, i.e. the selectivity of transformation to glycidol in relation to the allyl alcohol consumed, the conversion of allyl alcohol, the selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed. These functions were calculated according to the following formula:

Selectivity of transformation to glycidol, in relation to allyl alcohol consumed ($S_{\text{glyc/AA}}$):

$$S_{\text{glyc/AA}} = \frac{\text{amount of glycidol obtained (mol)}}{\text{amount of allyl alcohol consumed (mol)}} * 100\%$$

Conversion of allyl alcohol (C_{AA}):

$$C_{\text{AA}} = \frac{\text{amount of allyl alcohol consumed (mol)}}{\text{initial amount of allyl alcohol (mol)}} * 100\%$$

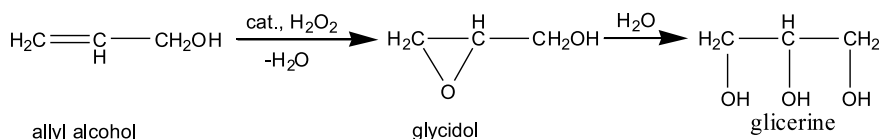
Selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed ($S_{\text{org/H}_2\text{O}_2}$):

$$S_{\text{org/H}_2\text{O}_2} = \frac{\text{amount of organic comp. obtained (mol)}}{\text{amount of H}_2\text{O}_2 \text{ consumed (mol)}} * 100\%$$

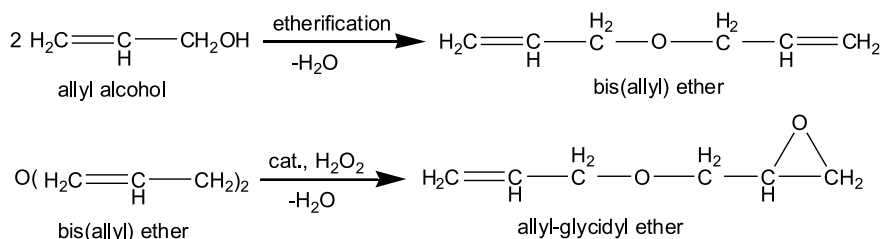
The conversion of hydrogen peroxide was calculated in a similar way as the conversion of allyl alcohol. The selectivity of transformations to by-products: glycerine, bis(allyl) ether and allyl-glycidyl ether were calculated in a similar way as the selectivity of transformation to glycidol.

RESULTS AND DISCUSSION

Epoxidation of allyl alcohol over titanium-silicalite catalysts by a 30 wt% water solution of hydrogen peroxide in methanol as a solvent leads to glycidol (2,3-epoxy-1-propanol) as the main product, which in the reaction conditions is partly hydrated to glycerine.



Depending on the conditions of epoxidation, also small amounts of ethers (bis(allyl) ether and allyl-glycidyl ether) are formed.



The influence of the titanium-silicalite catalyst concentration on the epoxidation of allyl alcohol over TS-1, TS-2 and Ti-MWW catalysts was tested in the range of its variation from 0.1 to 5.0 wt%. The obtained results present Figs. 6 – 8 (the changes of the following process functions: the selectivity of transformation to glycidol in relation to consumed allyl alcohol, the conversion of allyl alcohol, the selectivity of transformation to organic compounds in relation to consumed hydrogen peroxide). The conversion of hydrogen peroxide amounted to over 95 mol% and was not presented.

Figure 6 shows that in the case of TS-1 and TS-2 catalysts the best catalyst concentration is 1 wt% in order to achieve the highest selectivity to glycidol. The increase in these catalysts concentration causes the hydrolysis of epoxide ring and glycerine formation. The titanium silicalite catalysts take part in the hydrolysis of glycidol through the bonding of glycidol in active centers and polarization of epoxide ring. When the process is carried out in the presence of Ti-MWW catalyst the selectivity of the transformation to glycidol is constant in the whole catalyst concentration range and equals to 42 – 44 mol%.

The conversion of allyl alcohol increases with the increase of the TS-1 and Ti-MWW catalysts concentration, for TS-1 from 40 mol% to 64 mol% and for Ti-MWW from 74 mol% to 84 mol% – Fig. 7. In the range of low Ti-MWW catalyst concentration (0.1 – 1.0 wt%) the conversion of allyl alcohol increases to 90 wt% and then slightly decreases to 80 wt%, which is probably related to the decomposition of hydrogen peroxide over the catalyst. The results obtained for TS-2 and Ti-MWW catalysts are comparable and highly beneficial for the epoxidation process – mainly because the amount of allyl alcohol recycled for the process decreases.

Figure 8 shows that for the TS-1 catalyst concentrations in the range of 0.1 – 1.0 wt% the selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed rises from 63 mol% to 100 mol% and for the higher catalyst concentrations it does not change. It shows that for the TS-1 catalysts concentration in the range of 1.0 – 5.0 wt% the whole hydrogen peroxide introduced to the reactor takes part in the epoxidation process and does not decompose to oxygen and water. The epoxidation over Ti-MWW catalyst for the catalyst concentration 2.0 – 5.0 wt% shows the same course, but the

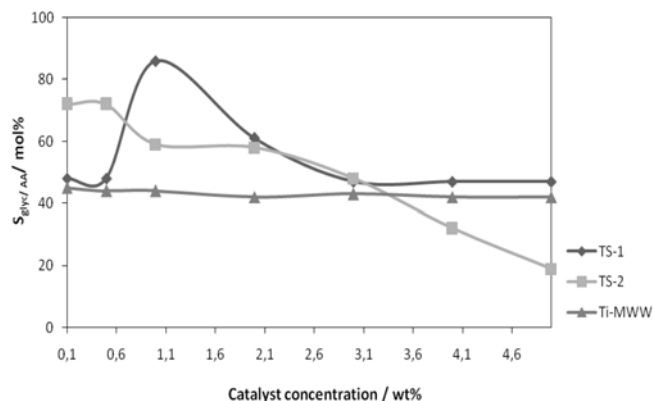


Figure 6. The influence of catalyst concentration on the selectivity of transformation to glycidol in relation to allyl alcohol consumed in the epoxidation over titanium-silicalite TS-1, TS-2 and Ti-MWW catalysts

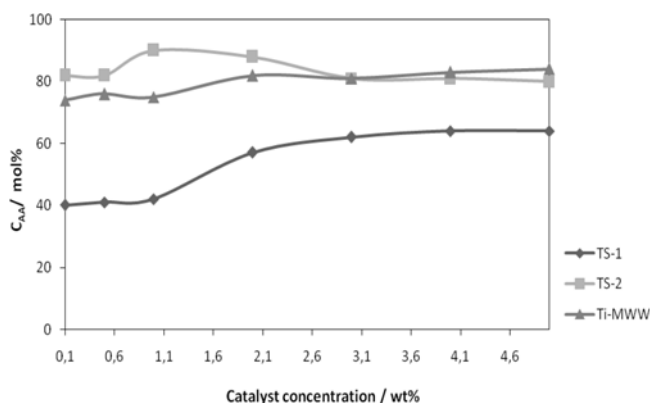


Figure 7. The influence of catalyst concentration on the conversion of allyl alcohol in the epoxidation over titanium-silicalite TS-1, TS-2 and Ti-MWW catalysts

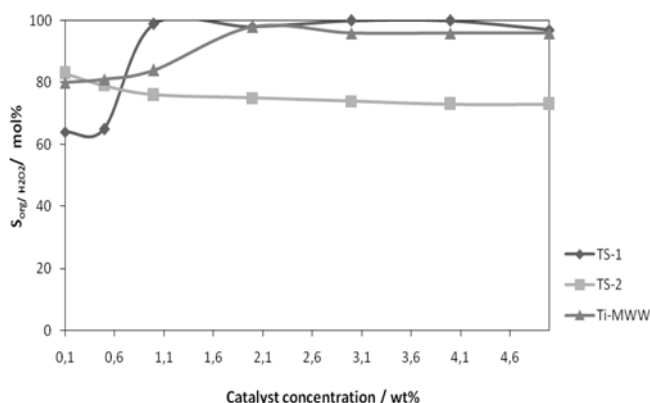


Figure 8. The influence of catalyst concentration on the selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed in the epoxidation over titanium-silicalite TS-1, TS-2 and Ti-MWW catalysts

values of the selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed are slightly lower (about 96 mol%). The selectivity of transformation to organic compounds in relation to hydrogen peroxide consumed for the TS-2 catalyst is almost constant and amounts to 72 – 82 mol%.

CONCLUSIONS

The presented results showed that the Ti-MWW catalyst can be applied as a catalyst in the epoxidation of allyl alcohol with hydrogen peroxide. The selectivity of transformation to glycidol for Ti-MWW catalyst is lower than for classical titanium silicalite TS-1 and TS-2 catalysts. Ti-MWW is more hydrophilic than TS-1 and TS-2 due to the silanols on the exterior surface and on the defect sites formed between the layers. It causes greater activity in the reaction of opening the epoxide ring in glycidol. For this reason the selectivity of transformation to glycidol in relation to allyl alcohol is lower than on TS-1 and TS-2, especially in the region of the low concentration of the examined catalysts – 0.1 – 1.0 wt%. The selectivity of transformation to organic compounds in relation to hydrogen peroxide for Ti-MWW catalyst is comparable with TS-1. A little lower is in the presence of TS-2. Conversion of allyl alcohol over the Ti-MWW catalyst is similar for TS-2 in the examined region of concentrations of catalysts (0.1 – 5.0 wt%). Both conversions are higher than on TS-1.

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