

Structural, optical, thermal and NLO behavior of zinc hydrogen maleate dihydrate single crystal

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Single crystal of zinc hydrogen maleate dihydrate (ZHMD) is grown by slow evaporation method at room temperature. The compound crystallizes in triclinic system with noncentrosymmetric space group P1. FT-IR and FT-Raman spectra of ZHMD are recorded. The versatile DFT is employed to understand the band structure of the crystal. Dipole moment of the molecule, ionization energy and electron affinity are established. The potential energy distribution PED of the vibrational modes is calculated using VEDA4. The results of the optimized molecular structure are compared with the single crystal XRD data. The Mulliken and NBO charges are calculated and interpreted. Optical properties of the grown crystal are analyzed using UV-Vis spectrum. Energy gap determined by Tauc plot is in a good agreement with the theoretical value calculated from HOMO-LUMO energies. Thermogravimetric analysis is done to assess the thermal behavior of the ZHMD crystal. Various thermodynamic parameters are calculated using the basis set B3LYP. The first order hyperpolarizability values of ZHMD establish its NLO nature. The molecular electrostatic potential MEP of the molecule is obtained. Third-order nonlinear response is studied using Z-scan technique and the corresponding parameters such as absorption coefficient, refractive index and third-order susceptibility are determined.

Keywords: ZHMD; UV-Vis; Tauc plot; DFT, enthalpy; Z-scan

1. Introduction

In the recent past, considerable attempt has been made to combine a metal ion with interesting organic and inorganic matrices to produce NLO materials with desirable properties. A metal-organic compound consists of metal or metal cluster centers that are connected by multidentate organic ligands. The interest in this class of compounds stems not only from the structural diversity they display but predominately due to their potential applicability as NLO materials. Zinc based single crystals possess a wide range of applications in the field of telecommunication and optical information storage devices. Zinc acetate dihydrate belongs to the family of inorganic nonlinear optical crystal. Maleic acid with relatively large

π -conjugation has attracted attention of crystal growers. Maleic acid forms crystalline maleate of various organic molecules through hydrogen bonding and π - π interactions and the intramolecular hydrogen bond in maleic acid is very strong [1, 2]. Maleic acid acts not only as an acceptor but also as an acidic ligand. It is reported that compounds of maleate such as L-arginine maleate dehydrate, L-phenylalaninium maleate, L-alaninium maleate and 2 aminopyridinium maleate dihydrate have improved NLO efficiency [3–5]. Maleates are of practical importance because they are efficient catalysts and also have medical significance. Recently, synthesis of ZHMD and its properties have been reported [6]. The hydrogen maleate ion has been studied by ab initio method [7]. In the present work, the influence of zinc on electron correlation and thereby hydrogen bonding in the crystal is investigated by DFT calculations. The results are reported for the first time.

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2. Experimental

ZHMD is synthesized by taking analytical grade chemicals of zinc acetate dihydrate and maleic acid in 1:1 ratio. The chemicals are dissolved in double distilled water. The prepared mixture is stirred well for 10 hours using a motorized magnetic stirrer at room temperature and a clear solution is obtained. The synthesized material is purified by repeated recrystallization process. Slow evaporation of the solvent yields small transparent, defect free crystals, optically clear and perfectly shaped. The crystals are obtained in a growth period of 6 days. The dimensions of the grown crystal are 0.5 cm \times 0.5 cm \times 0.2 cm. The photograph of the grown crystal is shown in Fig. 1. The grown ZHMD crystal is subjected to various characterization studies. FT-IR and FT-Raman spectra of ZHMD are recorded in the range of 4000 cm^{-1} to 400 cm^{-1} on Bruker IFS 66V spectrometer using KBr pellet technique. A Bruker kappa ApexII single crystal X-ray diffractometer with $\text{MoK}\alpha$ ($\lambda = 0.71073 \text{ \AA}$) radiation is used to obtain the structure of ZHMD crystal. The optical transmission spectrum of the crystal is recorded in the region of 200 nm to 800 nm, using a Varian/Cary 5000 UV-Vis-NIR spectrophotometer. The thermal behavior of the crystal has been investigated using PerkinElmer thermal analyzer at Department of Chemistry, IIT Madras. The third order non-linearity of the grown crystal is studied using the Z-scan technique. The studies have been carried out at NIT Trichy.

3. Computational details

The entire quantum chemical calculations are performed by DFT (B3LYP) method of 6-311++G(d,p) using the Gaussian 03W program. No imaginary frequencies are observed assuming that optimized geometry is obtained. The bond length and bond angles obtained from single crystal XRD are used to determine reduced masses, force constants, infrared intensities band assignment, Raman activities and depolarization ratio [8]. The vibrational modes are confirmed using Vibrational Energy Distribution Analysis VEDA program [9].



Fig. 1. Photograph of the as grown ZHMD crystal.

The assignments of the calculated wavenumbers are made on the basis of the Gaussview program [10]. On the basis of the optimized ground state geometries, TD-DFT calculation is performed at the same level of accuracy to investigate the absorption and electronic properties of ZHMD crystal. The excitation energy and oscillator strength are discussed using 6-311++G (d,p) basis set [11, 12]. The natural bond orbital (NBO) calculation [13, 14] is performed using NBO3.1 program contained in the Gaussian W03 package in order to understand various second order interactions which are a measure of intermolecular and intramolecular delocalizations.

4. Results and discussion

4.1. Molecular geometry

The crystal data obtained from single crystal XRD are given in Table 1. The optimized structural parameters of ZHMD calculated by B3LYP method with 6-311++G(d,p) basis set are listed in Table 2. The optimized structure is compared with experimental data. The ORTEP and optimized molecular structure of the title molecule is given in Fig. 2 and Fig. 3, respectively. The theoretical values of bond lengths and bond

angles calculated using 6-311++G(d,p) basis set are found to slightly differ with respect to the experimental values obtained for single crystal XRD. This is due to the fact that probably the theoretical calculations are done assuming an isolated molecule in gaseous phase whereas the experimental results apply to a molecule in solid state [15]. From Table 2 it is seen that the bond O1–Zn has a value of 2.1311 Å which is maximum, while the bond O5–H5 with a value of 0.79815 Å is the least. Thus, it is observed that the bond lengths calculated by B3LYP/6-311++G(d,p) basis set agree well with the experimental values [16–18].

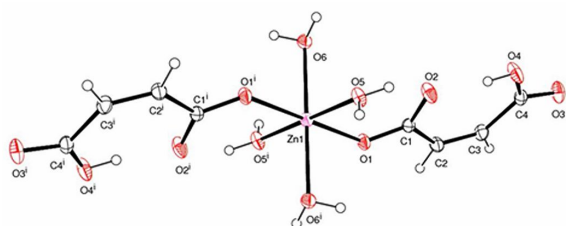


Fig. 2. ORTEP diagram of ZHMD single crystal.

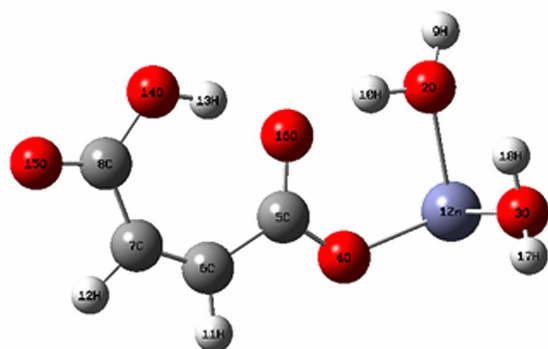


Fig. 3. Optimized molecular structure and atomic numbering of ZHMD single crystal.

4.2. Vibrational assignments

The FT-IR spectrum of a compound results due to superposition of absorption bands of specific functional groups. The vibrational band assignment of the molecule is made by observing the position, shape and relative intensities of the bands in FT-IR spectrum. The assignments are summarized

in Table 3. The O–H, C–H and C–O stretching vibrations are characteristic with regard to their shape and frequency and generally are observed in the region of 3600 cm^{-1} to 1700 cm^{-1} . The first feature observed in the vibrational spectrum of ZHMD is the broad band at 3386 cm^{-1} . It is the stretching vibration of the O–H bond present in the molecule. The O–H stretching vibration generally appears as a broad band extending from 3250 cm^{-1} to about 3550 cm^{-1} . Its appearance in this region suggests that the O–H stretching vibration is influenced by hydrogen bonding [19, 20]. The ring stretching band appears at 1505 cm^{-1} [21, 22]. The absorption band at 1193 cm^{-1} belongs to the asymmetric stretching vibration of the C–O group. The bending vibrations are generally found at lower wave numbers. The bands observed at 714 cm^{-1} and 573 cm^{-1} are assigned to O–C–C, C–O–H and C–C–C bending of the ring [23–25]. The bands at 566 cm^{-1} to 250 cm^{-1} in the FT-Raman spectrum are due to C–H out of plane deformation. The ZHMD molecule consists of 18 atoms, therefore, they have 48 normal modes. All the 48 fundamental vibrations are active in both IR and Raman spectra. The harmonic vibrational frequencies calculated for ZHMD by the B3LYP method using the 6-311++G(d,p) basis set are listed in Table 3 along with the frequencies as observed in the FT-IR spectrum. The experimental and theoretical FT-IR and FT-Raman spectra of ZHMD are shown in Fig. 4 and Fig. 5.

4.3. UV-Vis studies

UV-Vis spectroscopy correlates the electronic excitations between the energy levels that correspond to the molecular orbitals of the system [26–28]. UV-Vis spectroscopy is mostly used to identify conjugated systems which tend to have a stronger absorption. The recorded UV-Vis spectrum is shown in Fig. 6. The spectrum shows a lower cut off wavelength at 310 nm which corresponds well with the theoretical value of 324.45 nm computed using B3LYP basis given in Table 4. Fig. 7 presents the pictorial illustration of the frontier molecular orbital and corresponding positive and negative regions. In HOMO, the positive region

Table 1. Crystal data and structure refinement of ZHMD.

Crystal data	Data collection
Empirical formula	C ₄ H ₇ O ₆ Zn
Formula weight	367.56
Temperature	293(2) K
Wavelength	0.71073 Å
Crystal system, space group	Triclinic, P-1
Unit cell dimensions	a = 5.2153(2) Å alpha = 108.7240(10) deg. b = 7.3036(2) Å beta = 104.6410(10) deg. c = 9.2014(3) Å gamma = 93.1950(10) deg.
Volume	317.555(18) Å ³
Z	1
Calculated density	1.922 Mg/m ³
Absorption coefficient	2.000 mm ⁻¹
F(000)	188
Crystal size	0.35 mm × 0.30 mm × 0.25 mm
Theta range for data collection	2.44° to 26.00°
Limiting indices	-6 ≤ h ≤ 6, -8 ≤ k ≤ 8, -11 ≤ l ≤ 11
Reflections collected/unique	9670/1251 [R(int) = 0.0183]
Max. and min. transmission	0.6397 and 0.5410
Refinement method	Full-matrix least-squares on F ²
Goodness-of-fit on F ²	1.109
Final R indices [I > 2sigma(I)]	R1 = 0.0148, wR2 = 0.0423
R indices (all data)	R1 = 0.0148, wR2 = 0.0423
Extinction coefficient	0.089 (6)
Largest diff. peak and hole	0.266 e·Å ⁻³ and -0.169 e·Å ⁻³

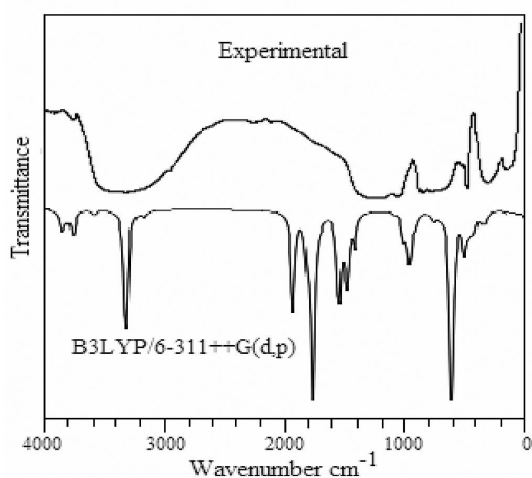


Fig. 4. Experimental and theoretical FT-IR spectra of ZHMD.

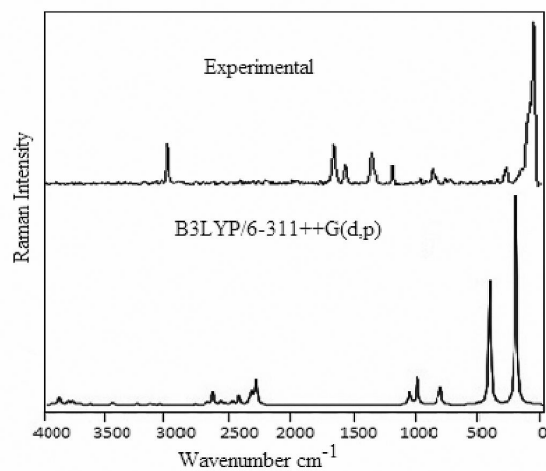


Fig. 5. Experimental and theoretical FT-Raman spectra of ZHMD.

Table 2. Geometrical parameters optimized in ZHMD by B3LYP with 6-31++G(d,p) basis set.

Parameter	B3LYP/6-31++G(d,p)	Experimental
Bond length		
C1–O1	1.264	1.251
C1–O2	1.274	1.26
C1–C2	1.481	1.489
C2–C3	1.347	1.332
C2–H2	1.082	0.93
C3–C4	1.496	1.487
C3–H3	1.083	0.93
C4–O3	1.238	1.222
C4–O4	1.352	1.288
O1–Zn	2.058	2.131
O5–H5	0.866	0.798
O4–H4	1.007	0.82
O5–Zn	2.156	2.035
O6–Zn	2.102	2.12
Bond angle		
[°]		
O1–C1–O2	122.5	123
O1–C1–C2	116.7	116.4
O2–C1–C2	120.8	120.6
C3–C2–C1	130.8	130.1
C3–C2–H2	117.5	115
C1–C2–H2	116.7	115
C2–C3–C4	132.5	130.7
C2–C3–H3	114.8	114.6
C4–C3–H3	112.7	114.6
O3–C4–O4	121.3	121.4
O3–C4–C3	118.4	118.5
O4–C4–C3	120.2	120.1
C1–O1–Zn	130.1	127.2
C4–O4–H4	110.7	109.5
Zn–O5–H5A	121.8	120
Zn–O5–H5B	117	119
Zn–O6–H6A	118.2	117.6
Zn–O6–H6B	105.8	106.4

is over Zn, while the negative sites are partially over O3. Positive and negative sites are spread over O2, O4, C5, C6 and O16. In LUMO, the negative and positive sites are observed over O14, O15, O4, while the groups of (C7, C8, H12) and

(C5, C6, H11) are partially covered with positive and negative sites. The more highly conjugated the system, the smaller HOMO-LUMO gap [29, 30]. Gauss-Sum 2.2 Program [31] is used to calculate group contributions to the molecular orbitals and prepare the density of the state DOS as shown in Fig. 8. The DOS spectrum is created by convoluting the molecular orbital information with the gaussian contour of the unit height. To ascertain the optical properties of the crystal, the values of optical constants have been determined from the recorded UV-Vis absorption spectrum as: extinction coefficient $K = 65.52$, reflectance $R = 1.98$ and refractive index $n = 1.3$, computed using standard formula [32]. The energy gap calculated by the B3LYP method is found to be 3.2613 eV which agrees with the energy gap calculated by the Tauc model, i.e. 3.74 eV as shown in Fig. 9 [33].

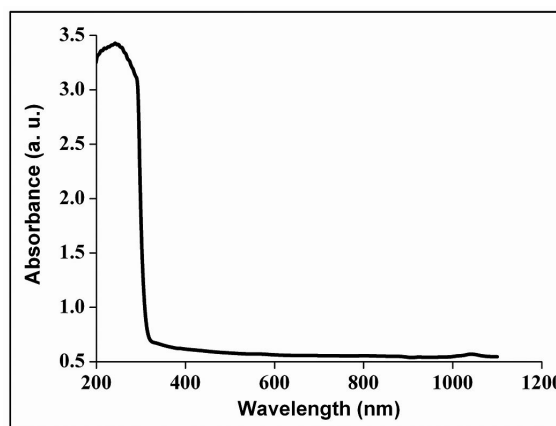


Fig. 6. UV-Vis spectrum of ZHMD crystal.

4.4. Mulliken and Natural Population Analysis

Population analysis is an investigation of charge distribution within molecules. Mulliken population analysis MPA is established as the linear combination of atomic orbitals and therefore the wave function of the molecule [34]. The Natural Population Analysis NPA is different from conventional MPA and seems to exhibit improved numerical stability. It illustrates the electron distribution in compounds. NPA analysis has been carried out using B3LYP basis, and the results are compared with

Table 3. Detailed assignment of observed and calculated vibrational frequencies of ZHMD along with potential energy distribution.

Experimental frequency [cm^{-1}]		Calculated wavenumber [cm^{-1}] B3LYP 6/-311++G(d,p)		Vibrational assignments with PED %
FT-IR	FT-Raman	FT-IR	Veda	
3773		3762	3762	OH ν (99 %)
3386	3245	3226	3225	OH ν (92 %)
	3053	3078	3077	OH ν (92 %)
	2462	1800	1800	OC ν (78 %)
	2257	1680	1681	CC ν (20 %) + HOH b (54 %)
	2039	1678	1678	CC ν (51 %) + HOH b (17 %)
	1812	1647	1647	HOH b (73 %)
	1703	1594	1594	OC ν (82 %)
1505	1609	1504	1504	HOC b (74 %)
	1397	1433	1433	OC ν (11 %) + CC ν (34 %) + HCC b (36 %)
	1227	1368	1368	OC ν (24 %) + HCC b (55 %)
1193		1222	1222	OC ν (23 %) + HCC b (59 %)
	1001	1042	1042	HCCC t (37 %) + HCCO t (52 %)
	899	891	891	HOCC t (65 %) + OCOC out (11 %)
863		867	867	OC ν (15%) + CC ν (14 %) + OCC b (27 %)
	800	803	803	CC ν (16%) + OCO b (23 %) + OCC b (10 %) + CCO b (10 %)
714		798	798	CCCC t (16 %) + OCOC out (52 %)
573		575	505	HOH b (13 %) + HOZnO t (68 %)
	377	394	358	CCC b (27 %) + COZn b (21 %) + OCC b (15 %)
	303	309	291	ZnO ν (11 %) + HOZnO t (21 %) + CCCC t (33 %)
	80	85	80	COZn b (25 %) + OZnO b (13 %) + CCOZn out (25 %)

ν = stretching, b = bending, γ = out of plane bending, t = torsion

that of Mulliken analysis. Corresponding pictorial charge distribution is given in Fig. 10.

4.5. Thermal studies

The TGA and DTA traces of ZHMD are shown in Fig. 11. To determine the stability of the grown crystal, TGA has been recorded at a scanning rate of 20 °C/min from 50 °C to 800 °C. It is seen that the crystal is thermally stable up to 150 °C. Decomposition of ZHMD occurs in three stages involving dehydration and decomposition. The first weight loss of 49.8 % is due to the decomposition of maleic acid, the second weight loss of 7.48 % occurs due to decomposition of zinc acetate and the third weight loss of 11.74 % is due to the residue.

4.6. Thermodynamic properties

The statistical thermochemical analysis of ZHMD is performed considering the molecule to be at room temperature and upon atmospheric pressure. Thermodynamic properties such as zero point energy, rotational constants, rotational temperatures, molar capacity, energy and entropy of ZHMD calculated using B3LYP/6-311++G(d,p) basis set are listed in Table 5. From the theoretical harmonic frequencies obtained by B3LYP method in the temperature range of 100 K to 1000 K, the thermodynamic functions such as entropy S , specific heat capacity at constant pressure C_p and enthalpy change ΔH for the title compound are evaluated and for example at 100 K the values are 663.74 J/mol·K, 218.98 J/mol·K

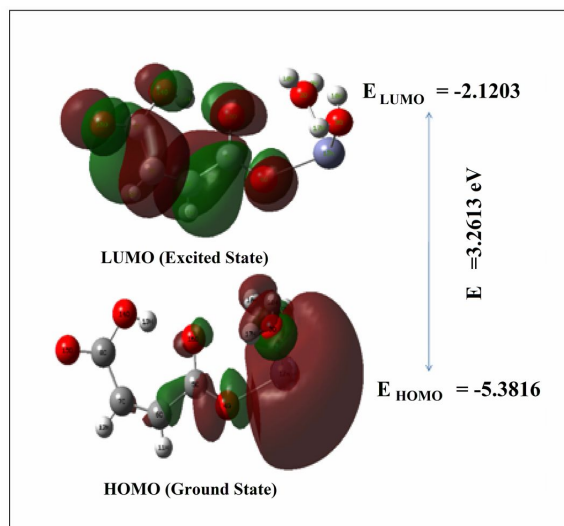


Fig. 7. Molecular orbitals and energies for the HOMO and LUMO of ZHMD.

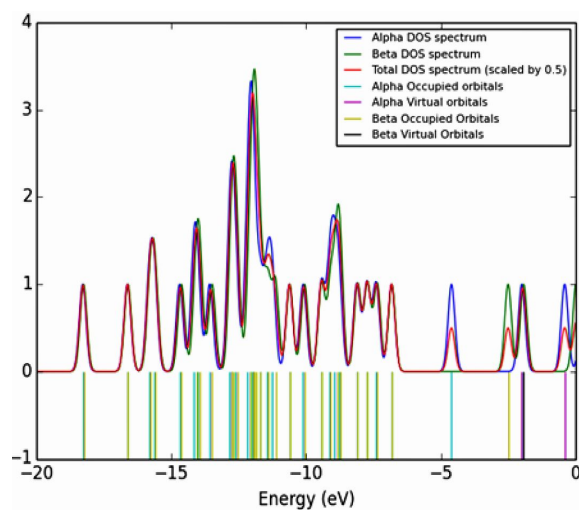


Fig. 8. DOS spectrum of ZHMD.

and 164.14 kJ/mol·K, respectively. With the increase in temperature, the thermodynamic properties are observed to increase due to the fact that the vibrational intensities of molecules increase with temperature [35, 36]. The best fit between the two, viz. thermodynamic properties and temperature is established by the quadratic formula and the corresponding correlation graphs are given in Fig. 12.

$$S = 262.84 + 0.6526T - 2.573 \times 10^{-4}T^2$$

$$R^2 = 0.9998$$

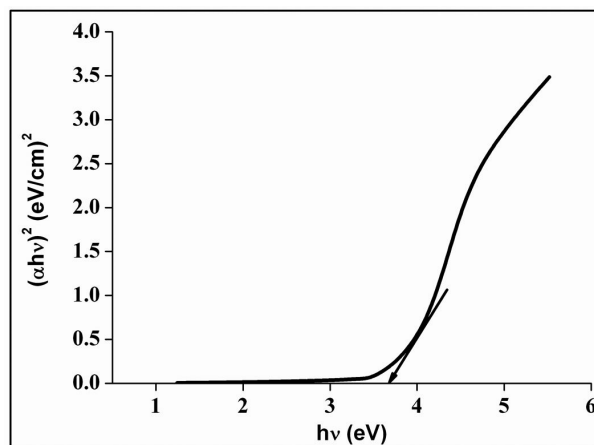


Fig. 9. Tauc plot of ZHMD crystal.

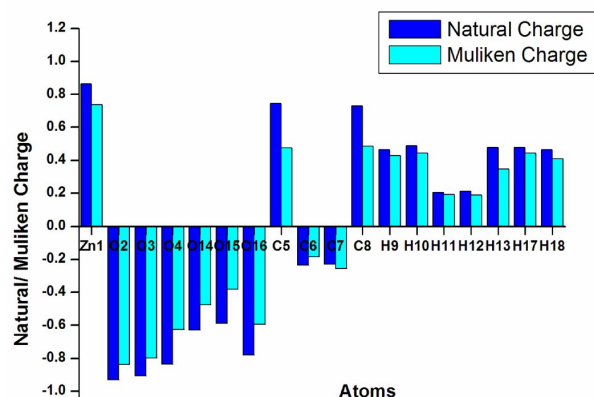


Fig. 10. Plot of Mulliken charges obtained by B3LYP/311++G(d,p) methods.

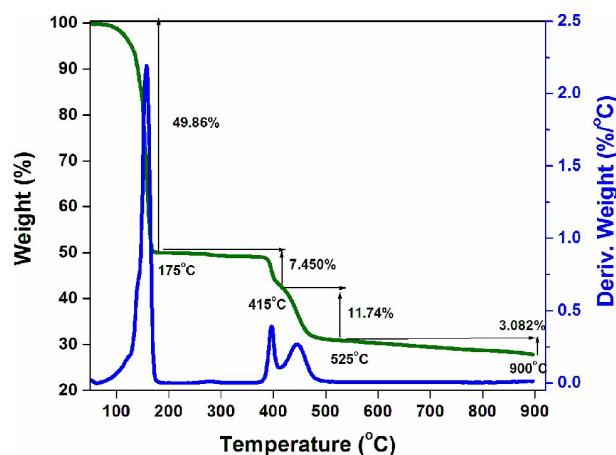


Fig. 11. TG-DTA graph of ZHMD crystal.

$$C_p = 59.34 + 0.331T - 1.7463 \times 10^{-4}T^2$$

$$R^2 = 0.9999$$

$$\Delta H = -7.433 + 0.1069T + 6.5 \times 10^{-5}T^2$$

$$R^2 = 0.9998$$

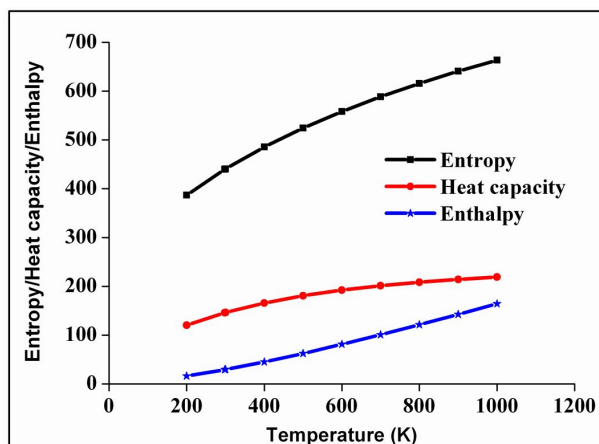


Fig. 12. Correlation graph between entropy, heat capacity and enthalpy with temperature.

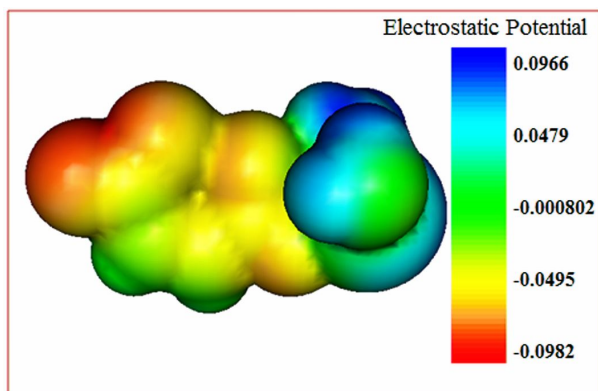


Fig. 13. The molecular electrostatic potential surface of ZHMD single crystal.

4.7. First order hyperpolarizability

Nonlinear optics NLO is the branch of optics that elucidates the behavior of light in nonlinear media. The dielectric polarization P behaves nonlinearly in the electric field E of the light. The second, third order and even higher harmonics

are obtained by multiplying the fundamental frequency with the induced polarization. The hyperpolarizability is a third rank tensor characterized by $3 \times 3 \times 3$ matrix. The 27 components of the 3D matrix can be reduced to 10 components by Kleinman symmetry [37]. The first order hyperpolarizability β of the molecule along with related properties calculated using B3LYP method is presented in Table 6. It follows from the table that the values of dipole moment and hyperpolarizability are greater than those of urea. Hence, it is concluded that the ZHMD single crystal has better nonlinear properties.

4.8. Natural Bond Orbital Analysis

NBO analysis is a useful tool to get an intuitive description of both electron orbital and population analysis [38]. Using the 6-311++G(d,p) basis, a second order perturbation theory study of Fock matrix in NBO basis is performed, which outlines the interaction between various molecular orbitals. Normally, these interactions are between bonding NBOs and non-bonding or antibonding orbitals. From the calculated values of E^2 , which correspond to the energy of hyper conjugative interaction, it is noted that for the interactive C8–O15 and C6–C7, the values are notable. The E^2 values are 29.03 kJ/mol and 37.22 kJ/mol, respectively. The larger the value of E^2 , the more intense is the interaction among electron donors and electron acceptors. Hence, there exists an intermolecular hyper conjugative interaction between $\Pi(C6-C7)$ and $\Pi(C8-O15)$. The ZHMD structure is stabilized by these interactions.

4.9. Molecular Electrostatic Potential

The molecular electrostatic potential MEP of a ZHMD molecule is a good guide to assess the molecule reactivity approaching positively or negatively charged reactants [39–41]. Mapping of the MEP on the molecular surface is performed with Gauss view shown in Fig. 13. It is seen in the MEP that the portions of positive potential are over all the hydrogen atoms (blue), while the negative potential sites are over the oxygen atoms (red). The green color represents zero electrostatic potential

Table 4. Experimental and calculated absorption wavelength λ , excitation state, oscillator strength [f], electronic absorption value [eV] and transition of ZHMD by TD-DFT method.

Excitation	Singlet A	Calc. λ [nm]	λ [nm]	Oscillator strength [f]	Electronic absorption [eV]	Transition
Excited state 1						
55 \leftrightarrow 56	0.99352	310	324.45	0.0064	2.6322	HOMO \leftrightarrow LUMO
Excited state 2						
49 \leftrightarrow 56	-0.22169		306.81	0.0003	3.308	HOMO-7 \leftrightarrow LUMO
51 \leftrightarrow 56	0.25802					HOMO-5 \leftrightarrow LUMO
52 \leftrightarrow 56	0.50289					HOMO-4 \leftrightarrow LUMO
53 \leftrightarrow 56	-0.33325					HOMO-3 \leftrightarrow LUMO
54 \leftrightarrow 56	-0.34778					HOMO-2 \leftrightarrow LUMO
Excited state 3						
55 \leftrightarrow 57	0.98053		300.52	0.0747	3.382	HOMO \leftrightarrow LUMO + 1

Table 5. Calculated thermodynamic parameters of ZHMD crystal.

Parameter	B3LYP/6-311++G(d,p)
Zero point vibrational energy [kcal/mol]	83.58763
Rotational constant [GHz]	2.10526
	0.34084
	0.34336
Rotational temperatures [K]	0.09672
	0.01636
	0.01552
Entropy [cal/mol·K]	105.111
Molar capacity at constant volume [cal/mol·K]	32.891
Energy [kcal/mol]	88.977

which is on the carbon atoms. The MEP, thus, simultaneously displays molecular size as well as positive, negative and neutral electrostatic regions of the title molecule. The colors seen in MEP confirm that intermolecular interaction can occur in the crystal.

4.10. Global reactivity descriptors

The overall stability of the system is assessed by measuring the global reactivity descriptors. The global electrophilicity index ω introduced by Parr [42, 43] is calculated using electronic potential μ and chemical hardness η . The values of ionization potential μ , chemical hardness η , softness S , electronegativity χ and electrophilicity

ratio ω are -3.7509 , 1.6306 , 0.3066 , 3.7509 and 4.314 , respectively. The value of μ indicates that the grown crystal is a soft molecule, hence, more easily polarizable.

5. Third order nonlinear optical properties

The Z-scan is a successful technique to study the third order NLO properties namely the nonlinear refractive index n_2 , nonlinear absorption coefficient β and the third order susceptibility χ^3 . Fig. 14 and Fig. 15, show the normalized transmission at the open and closed aperture. The transmission near the focus indicates

Table 6. Electric dipole μ , polarizability α , hyperpolarizability β of ZHMD crystal.

Parameter	esu ($\times 10^{-24}$)	Parameter	esu ($\times 10^{-33}$)
α_{xx}	14.1291	β_{xxx}	-501.0506
α_{xy}	-0.5216	β_{xxy}	78.6149
α_{yy}	14.480	β_{xyy}	-118.213
α_{xz}	3.0744	β_{yyy}	195.7825
α_{yz}	2.0548	β_{xxz}	-162.4525
α_{zz}	20.8813	β_{xyz}	57.0286
α_{tot}	16.4971	β_{yyz}	-80.5841
$\Delta\alpha$	25.3422	β_{xzz}	-277.7643
μ_x	-0.2471	β_{yzz}	37.5131
μ_y	-0.0018	β_{zzz}	-624.7585
μ_z	-0.4226	β_{tot}	1286.4736
μ	0.4895		

the nonlinear absorption, which is regarded as two-photon absorption. As seen from the closed aperture Z-scan curve, the prefocal transmittance peak is followed by the post focal valley which suggests that the origin of the nonlinear refractive index is thermo-optic. The nonlinear parameters are determined for ZHMD crystal using the standard relationships formulated by Sheik-Bahae *et al.* [44]. The nonlinear absorption coefficient $\beta = 0.5594 \times 10^{-3} \text{ cm/W}$ and the value of third order nonlinear susceptibility $\chi^3 = 3.0369 \times 10^{-7} \text{ esu}$. The nonlinear refractive index of the grown crystal is $-9.4799 \times 10^{-9} \text{ cm}^2/\text{W}$. The value is negative which indicates that the lensing effect is of defocusing nature. The defocusing nature and two-photon absorption coefficient are essential for the construction of intense laser beams and also useful in night vision devices.

6. Conclusions

A transparent, optical quality, colorless single crystal of ZHMD is grown using slow evaporation solution growth technique. A detailed vibrational assignment has been made for the ZHMD crystal using FT-IR and FT-Raman spectra. The observed and simulated spectra agree well with good frequency fit by the DFT (B3LYP) method.

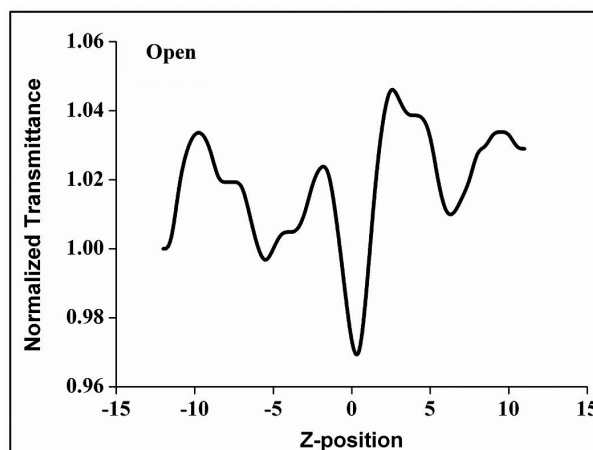


Fig. 14. Open aperture curve of ZHMD crystal.

The optical absorption in the entire visible region of the UV-Vis spectrum is flat and constant which suggests that the crystal is suitable for NLO applications. The optical parameters, such as extinction coefficient, reflectance and refractive index are determined. The HOMO-LUMO energy gap asserts the chemical reactivity of the molecule. Stability of the molecule proceeding from hyper-conjugative interaction and charge delocalization is confirmed using NBO analysis. NPA and MPA electron distribution in the compound is obtained. The MEP map of the crystal shows that the crystal has negative potential sites on oxygen and positive

potential sites around hydrogen. Thermodynamic parameters such as heat capacity, entropy and enthalpy are determined. From TGA it is seen that the crystal is thermally stable up to 150 °C, which suggests that it is suitable for device applications. The DFT calculations provide valuable insight into geometrical parameters, vibrational energies, rotational constant, thermodynamic properties, thermal energies, entropy, molar capacity constant and NLO activity of the compound. The calculated result obtained by DFT shows that the optimized geometry can well reproduce the crystal structure. The third order nonlinear properties of ZHMD crystal are confirmed by Z-scan technique, showing the crystal suitability for nonlinear optical devices.

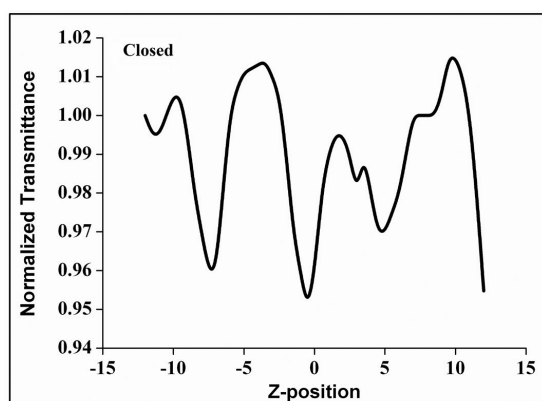


Fig. 15. Closed aperture curve of ZHMD crystal.

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