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Original paper

## Characteristics of sorbent products obtained by the alkaline activation of waste from waste incineration plants

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**Abstract.** In Poland, by 2020 430,000 Mg of hazardous waste will be formed annually by the combustion of waste. This waste must be properly managed so as not to endanger the environment. One promising way to manage selected waste is to process it in the synthesis of materials characterised by sorption properties. The results presented in this paper concern the possibility of producing sorbents from waste materials marked with codes 190112 and 190114, which came from two waste incineration plants in Poland. Alkaline activation was performed using two methods: a) hydrothermal, in a solution of 8 M NaOH at 75°C for 24 h; and b) in an autoclave, using a solution of 2 M NaOH at 140°C for 6 h. XRD analyses led to the identification of materials after synthesis of the following zeolite phases: analcime, chabazite and thomsonite. chabazite and analcime can be valuable absorbent materials.

**Key-words:** alkali-activated, ash and slag from waste incineration

## 1. Introduction

It is estimated that by 2020, approximately 430,000 Mg per year of hazardous waste will be produced in Poland (Wielgosiński, Naniecińska 2016). The residues from waste incineration are fly ash and slag, or ash - slag mixtures. It is important to properly manage them so that they do not pose a threat to the environment. The utilisation or the reprocessing of them is entered as part of the principles of the zero-waste economy, the effective use of resources (Resource Efficient Europe) and the closed-loop economy (Circular Economy). The idea of a Circular Economy is an economic model whose objective is to separate economic growth and development from the consumption of resources being depleted. "A circular economy is a systemic solution which relies on giving real priority to secondary raw materials, and even the obligation to use them, before we look to natural resources" (Łacka-Matusiewicz, Fraś 2012).

To correctly select the waste management and neutralisation method, above all its properties should be determined accurately. Most often, the following tests are performed: phase composition characterisation, chemical composition analysis, aqueous leaching testing, fineness and moisture content determination, morphology evaluation, pore size distribution and specific surface area measurements (the Brunauer–Emmett–Taller (BET) and Barrett–Joyner–Halenda (BJH) methods), as well as cation and anion exchange capacity determination (Łach et al. 2016). The results of the studies carried out allow the basic ways of using the waste to be defined.

One promising way to manage waste is to process it in the synthesis of zeolites or other materials, characterised by sorption properties (Charles et al. 2010). Zeolites are crystalline aluminosilicate compounds with a highly developed inner structure (Armbuster, Gunter 2001; Chica 2013) with a characteristic channel structure, in which the channels are interlinked with  $\text{SiO}_4$  and  $\text{AlO}_4$  tetrahedra. This makes them very good sorbents, catalysts and ion-exchange materials (Aiello 2002). Due to the high demand of the industry for synthetic zeolites, methods for producing these materials economically are still sought. In recent years, the feasibility of using fly ashes from hard coal and brown coal has been successfully studied (Chang, Shih 1998; Hollman et al. 1999; Inada et al. 2005; Derkowski 2006; Fotovat et al. 2009; Belviso et al. 2010; Franus et al. 2014) as well as ash from fluid boilers (Grela et al. 2016a, Grela et al. 2016b) for the synthesis of zeolite materials. Owing to intense studies, it is possible to discover new zeolite structures and to obtain thorough knowledge of their properties.

Ash from waste incinerators may be used as the starting material for the synthesis of zeolites such as gismondine or gmelite (Yang, Yang 1998; Miyake et al. 2002), zeolite A and zeolite P (Miyake et al. 2002; Tamura et al. 2006; Sallam et al. 2008). It is also possible to obtain zeolite materials from waste incineration slag (Chiang et al. 2012) in this way, e.g. tobermorite and hydrated sodium aluminium silicates were obtained that exhibited the BET specific surface area of  $22.1 \text{ m}^2 \text{ g}^{-1}$  (compared with  $4. \text{ m}^2 \text{ g}^{-1}$  for the starting material). The studies showed the suitability of the materials obtained for the adsorption of heavy metal cations. Zeolites produced from fly ash from waste incineration may be employed for the removal of cadmium, chromium and lead ions (Shim et al. 2003; Gupta et al. 2005; Tao 2006). Other investigations (Rodzewicz et al. 2016) have demonstrated that ash from

wastewater sludge incineration may be used as a filter filler for the treatment of phosphorus-containing wastewaters.

The aim of this study was to characterise and assess the physicochemical properties of the waste such as slag and ash from waste incineration, and to characterise the synthesis products formed in the autoclaving process.

## 2. Materials and methods

### 2.1. Starting materials

Two waste types with the EWC codes 190112 (bottom ash and slag other than those indicated by code 190111 - P1) and 190114 (fly ash other than those indicated by code 190113 - P2) from two waste incineration plants in Poland were subjected to study. The waste incineration in both the first and in the second plant is carried out in a rotary oven with an after-combustion chamber. Flue gases are cooled down in a recuperator boiler. Table 1 presents the designation of material samples used in this study.

TABLE 1

Designations of the materials used in the study.

Sample	Description
P1	190112 – bottom ash and slag other than those indicated by code 190111. Industrial and medical waste incineration plant.
P1A	P1 after synthesis in an autoclave, (2M NaOH, 140°C, 6h)
P1H	P1 after hydrothermal synthesis, (8M NaOH, 75°C, 24 h)
P2	190114 – fly ash other than those indicated by code 190113. The incineration of sewage sludge from waste water treatment plants.
P2A	P2 after synthesis in an autoclave, (2M NaOH, 140°C, 6h)
P2H	P1 after hydrothermal synthesis, (8M NaOH, 75°C, 24 h)

### 2.2. Synthesis

#### 2.2.1. Synthesis in autoclave

The synthesis process was carried out in a PARR pressurised chemical reactor (autoclave). The post-processing waste from the thermal conversion of the waste was denoted as P1 and P2. Bottom ash and slag containing dangerous substances was charged into the reactor with a 250 g / 1000 ml 2 M NaOH solution added. The mixture thus prepared was closed in the chemical reactor, in which the pressure was increased up to the level of 0.5 MPa. After 5 min, the mixture was heated up to 140°C at a rate of 2°C/min, while the working pressure in the reactor was concurrently raised. The heating time at the reaction temperature was equal to 6 h, and the working pressure was 1.1 MPa. After

heating, the reaction system was cooled down at a rate of  $0.3^{\circ}\text{C}/\text{min}$ . The total processing time was 24 h (Fig. 1).

Afterwards, the NaOH solution used for the synthesis was removed (by decanting the liquid from above the sediment) into an appropriate container. The obtained reaction mixture was placed in a vessel with water warmed up to  $30^{\circ}\text{C}$ . To wash the mixture, a POLSONIC ultrasonic washer was employed. The washing was carried out for 30 min. After the washings were poured out, the reaction mixture was washed with water until the pH of the washings fell below 9. Subsequently, the reaction mixture was dried at a temperature of  $100^{\circ}\text{C}$  for 12 h.

### 2.2.2. Hydrothermal synthesis

The respective measured amounts of waste P1 and P2 were suffused with 8 M NaOH solution and placed in polypropylene cylindrical vessels with a capacity of 1000 ml. The containers were sealed and placed in a drying oven at  $75^{\circ}\text{C}$  for 24 h. After 24 h, the final reaction mixture was cooled to room temperature and rinsed using the procedure described above (synthesis in autoclave).

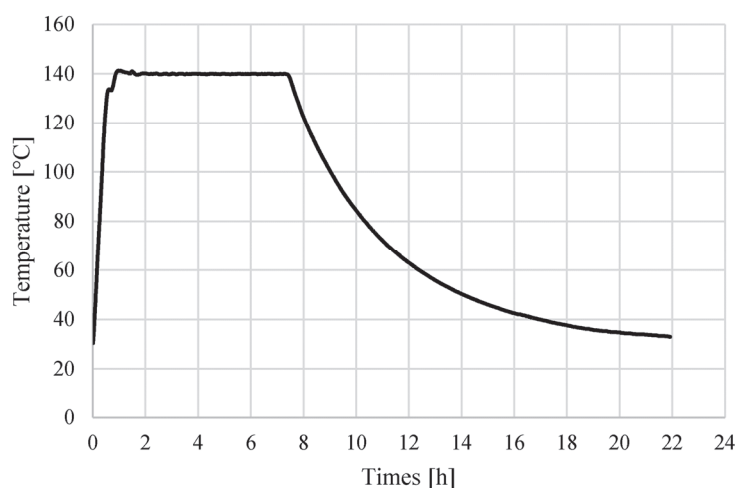


Fig. 1. Schematic diagram of the zeolite production process by autoclaving.

### 2.3. Examination methods

The phase composition of the samples – P1, P1A, P1N, P2, P2A and P2N – was determined using the powder X-ray diffraction (XRD) Debye–Scherrer–Hull method. Diffractograms of the samples were recorded with the use of a RigakuSmartLab X-ray diffractometer using the following parameters:  $\text{CuK}\alpha$  radiation, a graphite reflection monochromator, tube voltage 45 kV, tube current 200, step size  $2\theta = 0.05^{\circ}$ . The values of the interplanar distances were used for the identification of phases present in the samples

based on data contained in the International Centre for Diffraction Data (ICDD 2014) catalogue and XRAYAN software. For samples P1 and P2, their chemical composition was determined. The analyses were carried out using a XRF WDX PANalytical spectrometer.

The dioxin content was determined with gas chromatography coupled with mass spectrometry (GC-MS/MS) in the Laboratory of Trace Analysis at the Cracow University of Technology. The examination was carried out according to the P/01/03 procedure imposed on 11<sup>th</sup> March 2010. The results were expressed as the upper limit of toxic equivalency (TEQ) according to I-TEF (1988).

Aqueous leaching tests were performed in the Mo-BRUK S.A. accredited laboratory in accordance with PN-EN 12457-4:2006. Samples for the determination of metal content were prepared according to PN-EN ISO 15587-2:2005, depending on the type of substance analysed, and according to the standard code: gravimetry, spectrophotometry, ion-chromatography, or Inductively Coupled Plasma – Optical Emission Spectrometry (ICP-OES). The masses of the analysed samples were of ca. 2.5 kg each. Single-stages were, batch test, at a liquid to solid phase ratio of 10 l kg<sup>-1</sup> in the case of below 10 mm grain-size materials (without size reduction or with size reduction).

Fineness testing was carried out in accordance with EN 451-2 by wet sieving. Moisture content was determined in the delivered conditions by an analytical method of drying the ash in a moisture analyser to constant mass.

The morphology of samples was studied with a JEOL JSM-820 Scanning Electron Microscope (SEM). Samples were appropriately prepared beforehand. Small quantities of materials were dried to constant mass, and then were placed on a carbon substrate that ensured removal of the charge from the sample. The materials were coated with a thin layer of gold by a JEOL JEE-4X Vacuum Evaporator.

Porosity and specific surface area measurements were performed with the use of a BET-ASAP 2020 Accelerated Surface Area and Porosimetry Analyser. Samples were degassed at a temperature of 373 K for 24 h. To determine the distribution of pore volume function in the materials analysed, experimental low-temperature (77 K) nitrogen adsorption isotherms were used. The Barrett–Joyner–Halenda method was employed, and to determine the specific surface area, the Brunauer–Emmett–Taller method was applied.

The Cation Exchange Capacity (CEC) was determined with the NH<sub>4</sub> sorption–desorption method, while the Anion Exchange Capacity (AEC) was determined with the phosphate method (Grela, Bajda 2017).

### **3. Results and discussion**

#### **3.1. Raw materials**

Figures 2a and 2b present the recorded diffractograms of the waste samples, P1 and P2, respectively. Based on the analysis of the diffraction spectrum it was found that in sample P1 the following phases are present: quartz, perovskite, ilmenite, rutile, esseneite, olivine, feldspars, and other calcium and magnesium silicates, such as Ca<sub>2</sub>Mg<sub>6</sub>Si<sub>6</sub>O<sub>20</sub> (Fig. 2a); while in sample P2, the following phases were identified: quartz, calcium and aluminium phosphates, anhydrite, plagioclase, and wüstite (Fig. 1b).

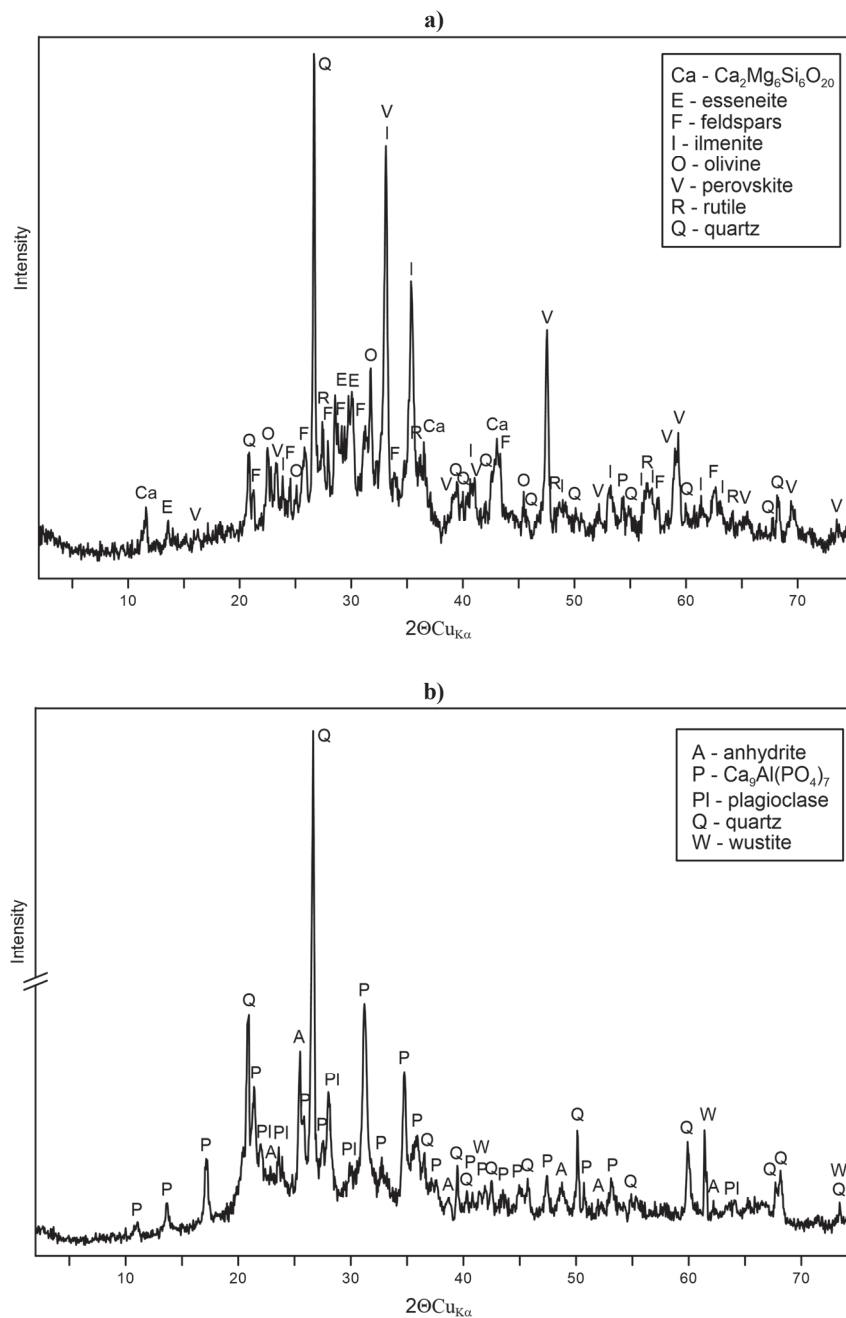


Fig. 2. X-ray diffraction pattern of samples: P1 (a):  $Ca_2Mg_6Si_6O_{20}$  PDF-2 No. 02-0558, Esseneite PDF-2 No. 25-0143, Feldspars PDF-2 No. 20-0572, Ilmenite PDF-2 No. 29-0733, Olivine PDF-2 No. 07-0164, Perovskite PDF-2 No. 42-0423, Rutile PDF-2 No. 34-0180, Quartz PDF-2 No. 05-0490; and P2 (b): Anhydrite PDF-2 No. 06-0226,  $Ca_9Al(PO_4)_7$  PDF-2 No. 48-1192, Plagioclase PDF-2 No. 41-1486, Quartz PDF-2 No. 05-0490, Wüstite PDF-2 No. 461312.

Table 2 presents the chemical composition analysis results for the waste samples. The dominating components are SiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub>, which are considered to play a key role in zeolite formation. CaO (ca. 16wt%) adversely affects the process of zeolite phase formation. The content of SiO<sub>2</sub> and Al<sub>2</sub>O<sub>3</sub> is higher in P2 (approx. 45%) in comparison with sample (P2), which contains 29%.

The analyses carried out of the dioxin content showed a total content of dioxins (I-PCDD/F-TEQ) (ng g<sup>-1</sup>): P1 = 0.30 ± 0.078 and P2 = 0.0092 ± 0.0024.

The result is expressed as standardised TEQ according to the I-TEF of 1988 and presented as its upper limit. The total dioxin content of P1 is almost thirty-two times higher than in P2.

Despite such differences, the dioxin content of both wastes is at a low level. In Poland, there are no regulations concerning dioxin and furan emissions from waste incineration, and their content in ashes and slags is not routinely determined.

TABLE. 2

The chemical composition (wt%) of samples P1 and P2 used in the study.

Sample	SO <sub>3</sub>	P <sub>2</sub> O <sub>5</sub>	SiO <sub>2</sub>	TiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>	Na <sub>2</sub> O	K <sub>2</sub> O	CaO	MgO	Cl	Br
P1	4.29	1.48	18.67	10.52	10.27	16.94	3.45	0.36	15.54	2.35	1.88	0.02
P2	2.75	21.69	20.04	0.71	24.84	5.15	1.01	1.11	16.09	2.84	0.03	0.00

Setting out dioxin emission standards would result in the necessity of carrying out appropriate controls. Dioxin emission testing is very costly and only a few research centres in Poland are able to take on the task of determining the dioxin concentration (Pająk 1996).

The toxicity of the ash and slag may be compared with the pollution limit values that are used for the ground or soil. A total dioxin content less than 0.005 ng g<sup>-1</sup> allows the area to be used for agricultural purposes without any limitations. If the total dioxin content is in the range of 0.04 - 0.1 ng g<sup>-1</sup>, then it is permitted to use the area for fruit and shrub growing (albeit with dioxin content control of other products), as well as for recreation grounds for children. A total dioxin content in the range of 0.1 - 1.0 ng g<sup>-1</sup> results in the total unsuitability of the material for use in agriculture and horticulture, and it must be replaced (Pająk 1996).

Table 3 presents the results of the aqueous leaching tests for P1 and P2. Based on the results obtained, it was concluded that high amount of molybdenum and sulphate ions are washed. Sample P1 contains an elevated amount of chloride ions. The elution of total dissolved solids (TDS) is twice as high for P1 (29.05 mg kg<sup>-1</sup>) compared to P2 (14.38 mg kg<sup>-1</sup>). Samples P1 and P2 have a similar pH in the range of 9.4.

Table 4 shows selected physical properties of the waste samples. The test results are significant for their technological importance. They affect the alkaline activation process with respect to their water demand. Water demand depends on the moisture content level and their grain size (fineness). The drier as well as the smaller the grains, the higher the absorptivity of the liquid alkaline activator. It was found that the studied waste samples differed slightly in their fineness, meant as the sieve residue (aperture size 0.045 μm) at wet

sieving. The results obtained show that the particle or grain size distribution of the tested waste samples are small.

Figures 3 shows the morphology of the waste samples. In both samples, the irregular shape of the grains and size variation can be observed. At high magnification, in sample P1 a plumose surface is partially visible, while in sample P2 the grains are smooth. Shapes of rods and blades can also be observed. In both cases, the grains of the waste often form agglomeration.

TABLE 3

Leaching tests results of samples P1 and P2 (mg kg<sup>-1</sup>).

Component	Permissible leaching limit values <sup>1)</sup> Liquid/solid phase = 10 l/kg [mg/kg dry substance]		190112	190114
	Criteria for the acceptance of the waste to be deposited at an inert waste landfill	Criteria for the acceptance of the waste to be deposited at a hazardous waste landfill	P1	P2
As	0.5	25	<0.5	<0.5
Ba	20	300	1.59	0.80
Cd	0.04	5	<1	<1
Cr	0.5	70	1.59	<1
Cu	2	100	<1	<1
Hg	0.01	2	<0.001	<0.001
Mo	0.5	30	10.73	4.87
Ni	0.4	40	<1	<1
Pb	0.5	50	<1	<1
Sb	0.06	5	<0.5	<0.5
Se	0.1	7	<0.5	<0.5
Zn	4	200	<1	<1
Cl <sup>-</sup>	800	25 000	7 737.4	<250
F <sup>-</sup>	10	500	7.52	52.2
SO <sub>4</sub> <sup>2-</sup>	1 000	50 000	7 511.2	8 313
DOC	500	1 000	<300	<300
TDS <sup>**)</sup>	4 000	100 000	29 050	14 380
Cr <sup>6+</sup>			<0.5	<0.5
pH			9.3	9.4

<sup>1)</sup> Permissible leaching limit values in the case of waste deposited at a landfill provided with leachate collection systems directed afterwards to the wastewater treatment plant, with the exception of DOC and TDS, are regarded as fulfilled in the case of values higher than set out in the Table.

<sup>2)</sup> Values for TDS may be used interchangeably for sulphates and chlorides.



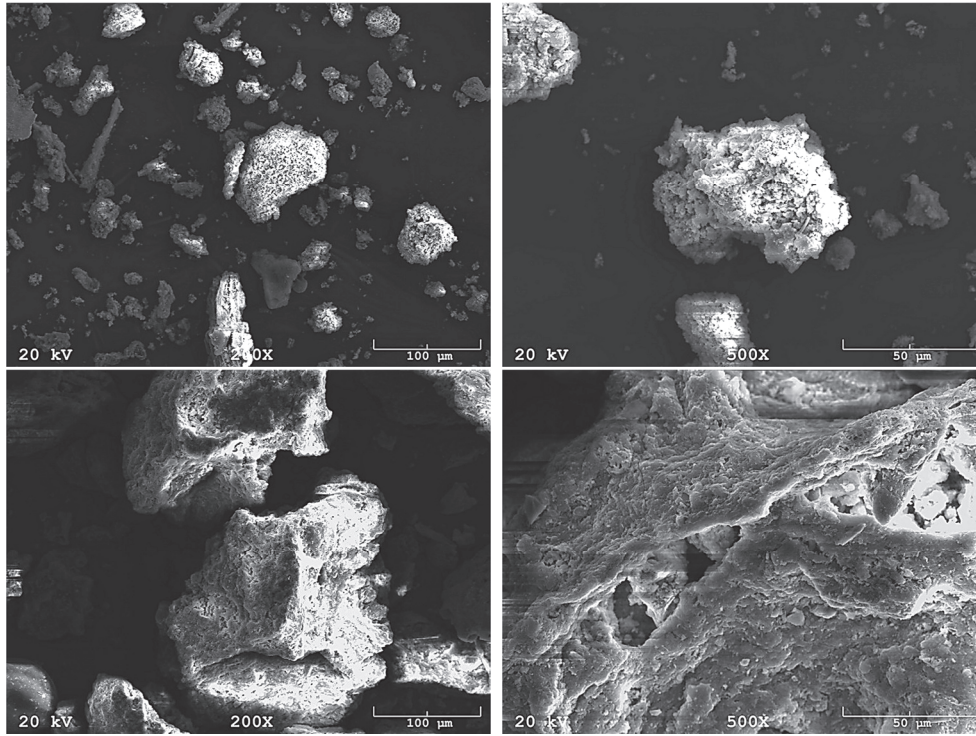


Fig. 3. SEM images of samples P1(a, b) and P2 (c, d).

TABLE 4

Textural testing results and CEC and AEC results for samples P1 and P2, and the results of the physical properties' testing of samples P1 and P2.

Sample	BET [m <sup>2</sup> /g]	BJH [m <sup>2</sup> /g]	Single point adsorption [cm <sup>3</sup> /g]	Adsorption average [nm]	BJH Adsorption [nm]	CEC [meq/100g]	AEC [meq/100g]	Fineness [%]	Moisture [%]
P1	20.85	16.35	0.035	6.80	8.63	5.58	34.07	87.88± 1.33	7.58± 0.09
P2	4.51	2.62	0.012	10.40	17.26	3.31	24.46	65.44± 0.70	0.23± 0.01

The textural parameters were determined for samples P1 and P2 (Table 4). Sample P1 has a type II isotherm, while sample P2 has a type III isotherm, but both have a hysteresis loop of type H3 (Fig. 4). This corresponds to a mesoporous character with the formation of slit-shaped pores arising from the stacking of crystal particles. The specific surface area values (SBET) of samples P1 and P2 were 20.85 m<sup>2</sup> g<sup>-1</sup> and 4.51 m<sup>2</sup> g<sup>-1</sup>, respectively. The total pore volume values were 0.035 cm<sup>3</sup> g<sup>-1</sup> and 0.012 cm<sup>3</sup> g<sup>-1</sup>, respectively (Table 4).

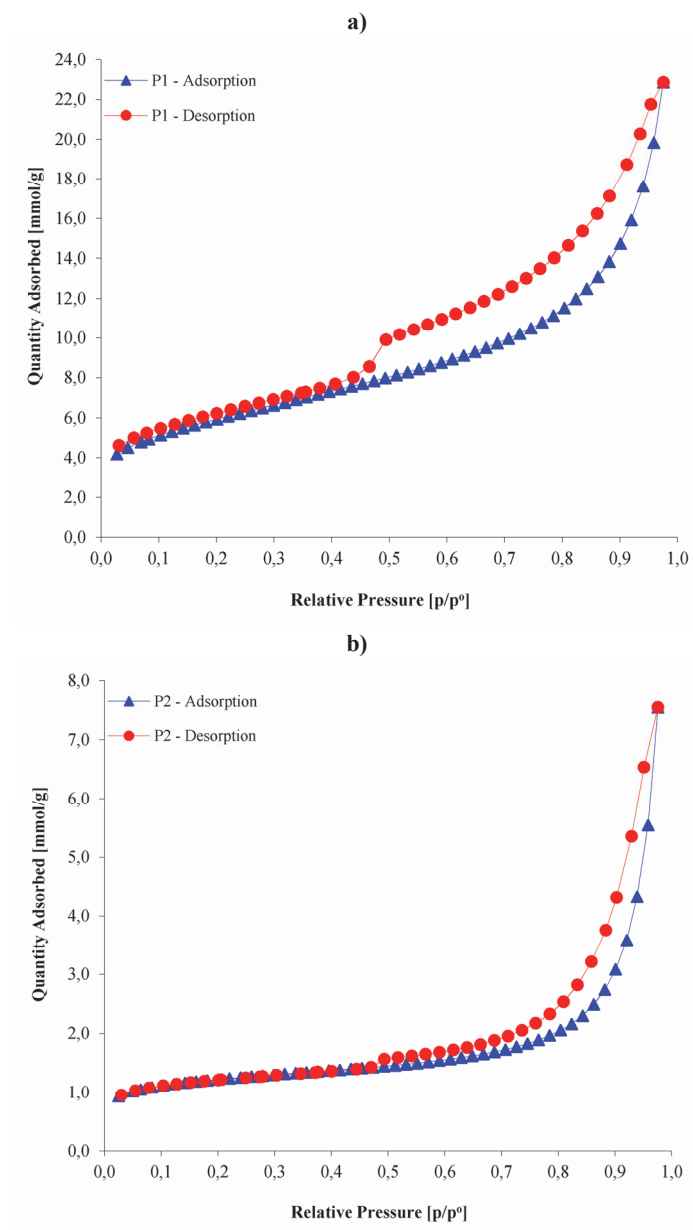


Fig. 4. BET adsorption and desorption isotherms: (a) P1, (b) P2.

The determined CEC and AEC values for samples P1 and P2 are listed in Table 4. The higher the ion-exchange capacity material, the higher the quantity of anions or cations retained on the material by the ion-exchange process. Comparing both samples, it can be determined that they have a much greater capacity for anion than cation exchange. Sample

P1 has a six times larger AEC than CEC. For sample P2, the AEC is eight times larger than the CEC.

### 3.2. Materials after synthesis

Figures 5a and 5b demonstrate the diffractograms of samples P1A (a) and P1H (b). XRD analysis allowed identification of the following in the studied materials. In sample P1A, quartz, perovskite, ilmenite, rutile and esseneite all occur in the starting material and the product of synthesis; chabazite from the zeolite groups; tobermorite, an alteration product of calcium minerals occurring in the substrate of synthesis; and hydrous sodium and titanium silicates, the product of the reaction between NaOH and Ti minerals. Whereas in sample P1H, the following phases are present: quartz, perovskite, ilmenite, rutile and unreacted reagents occurring in the starting samples; chabazite and thomsonite from the zeolite group; and tobermorite and hydrous sodium and titanium silicates.

Figures 6a and 6b present the recorded diffractograms of samples P2A and P2H, respectively, after hydrothermal synthesis. Based on the analysis of the diffraction spectrum, sample P2A contains two zeolites phases: analcime and thomsonite, and quartz as unreacted phase residue remaining in the matrix from the starting material. In sample P2H, the following phases were identified: unreacted quartz, zeolites phases or from the analcime and thomsonite group, and hydroxyapatite, a product of the transformation of initial phosphorous and calcium minerals in the presence of NaOH.

Synthetic zeolites (e.g. chabazite and analcime) obtained by the transformations of P1 and P2 can be a valuable absorbent material for the gaseous forms of  $\text{NH}_3$  and Hg, as well as  $\text{CO}_2$  and  $\text{SO}_2$  (Morency et al. 2002; Querol et al. 2002; Wdowin et al. 2012; Wdowin et al. 2014).

Figure 7 presents the morphology of the products obtained by alkali activation of the tested waste. There is a clear change in the morphological forms of the materials in comparison with the starting materials. In the case of sample P1H, a slightly feathered structure can be observed, the grains are made up of rods, and in most parts of interconnected plaques. These are structures with appearance characteristic of tobermorite and chabazite. Sample P1A has quite a different structure, the grains of which consist of hexagonal rods. Samples P2A and P2H consist of grains of spherical-surface-like structures, where stamens and plaques have crystallised.

Samples P1A and P1H show a similar type of isotherm and hysteresis loop of type H3, as in the case of the initial samples. However, we observed a relatively higher increase in the SBET value for sample P1 synthesised using the autoclave method ( $51.49 \text{ m}^2 \text{ g}^{-1}$ ) than in the case of the hydrothermal method ( $37.09 \text{ m}^2 \text{ g}^{-1}$ ). Interestingly, the total pore volume is higher for sample P1H ( $0.113 \text{ cm}^3 \text{ g}^{-1}$ ) than for P1A ( $0.071 \text{ cm}^3 \text{ g}^{-1}$ ). Samples P2A and P2H have similar SBET values ( $43.30 \text{ m}^2 \text{ g}^{-1}$  and  $46.26 \text{ m}^2 \text{ g}^{-1}$ , respectively), but the total pore volume of sample P2A ( $0.114 \text{ cm}^3 \text{ g}^{-1}$ ) is nearly twice the value of sample P2H ( $0.056 \text{ cm}^3 \text{ g}^{-1}$ ) (Table 6). This indicates that the synthesis method had a significant influence on the textural parameters. (Table 6).

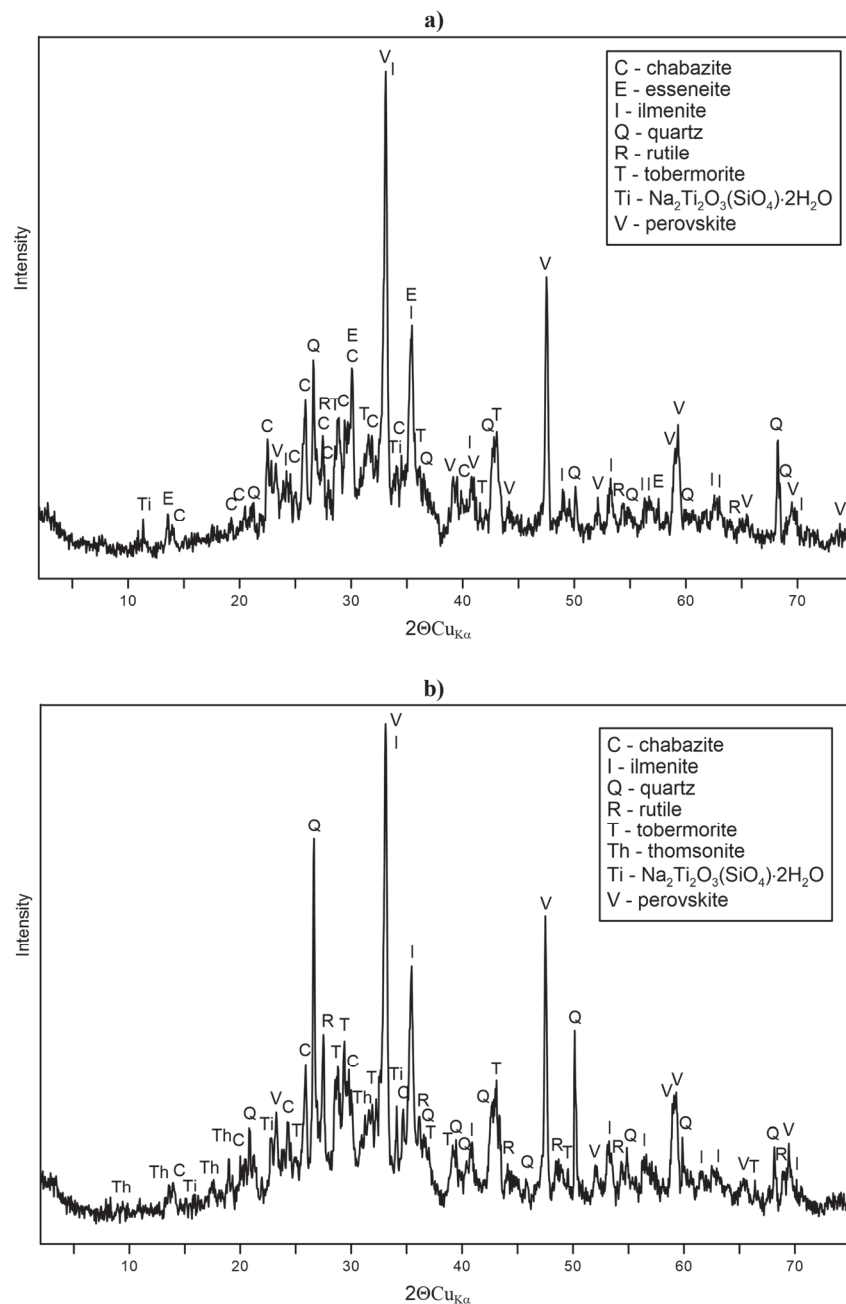


Fig. 5. X-ray diffraction pattern of samples. P1A (a): Chabazite PDF-2 No. 44-0248, Esseneite PDF-2 No. 25-0143, Ilmenite PDF-2 No. 29-0733, Quartz PDF-2 No. 05-0490, Rutile PDF-2 No. 34-0180, Tobermorite PDF-2 No. 10-0374,  $\text{Na}_2\text{Ti}_2\text{O}_3(\text{SiO}_4) \cdot 2\text{H}_2\text{O}$  PDF-2 No. 47-0591, Perovskite PDF-2 No. 42-0423; and P1H (b): Chabazite PDF-2 No. 44-0248, Ilmenite PDF-2 No. 29-0733, Quartz PDF-2 No. 05-0490, Rutile PDF-2 No. 34-0180, Tobermorite PDF-2 No. 10-0374, Thomsonite PDF No. 46-1448,  $\text{Na}_2\text{Ti}_2\text{O}_3(\text{SiO}_4) \cdot 2\text{H}_2\text{O}$  PDF-2 No. 47-0591, Perovskite PDF-2 No. 42-0423.

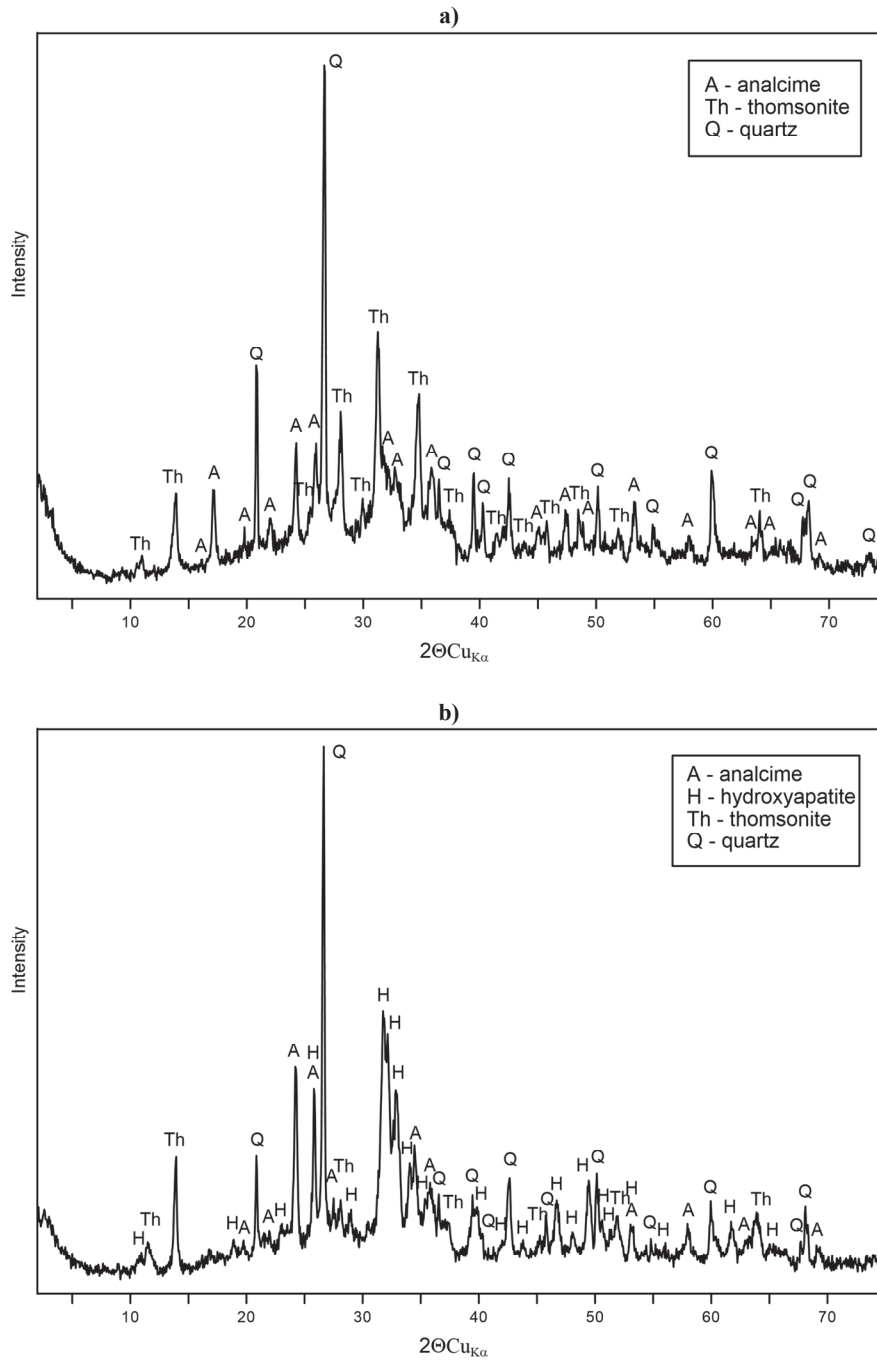
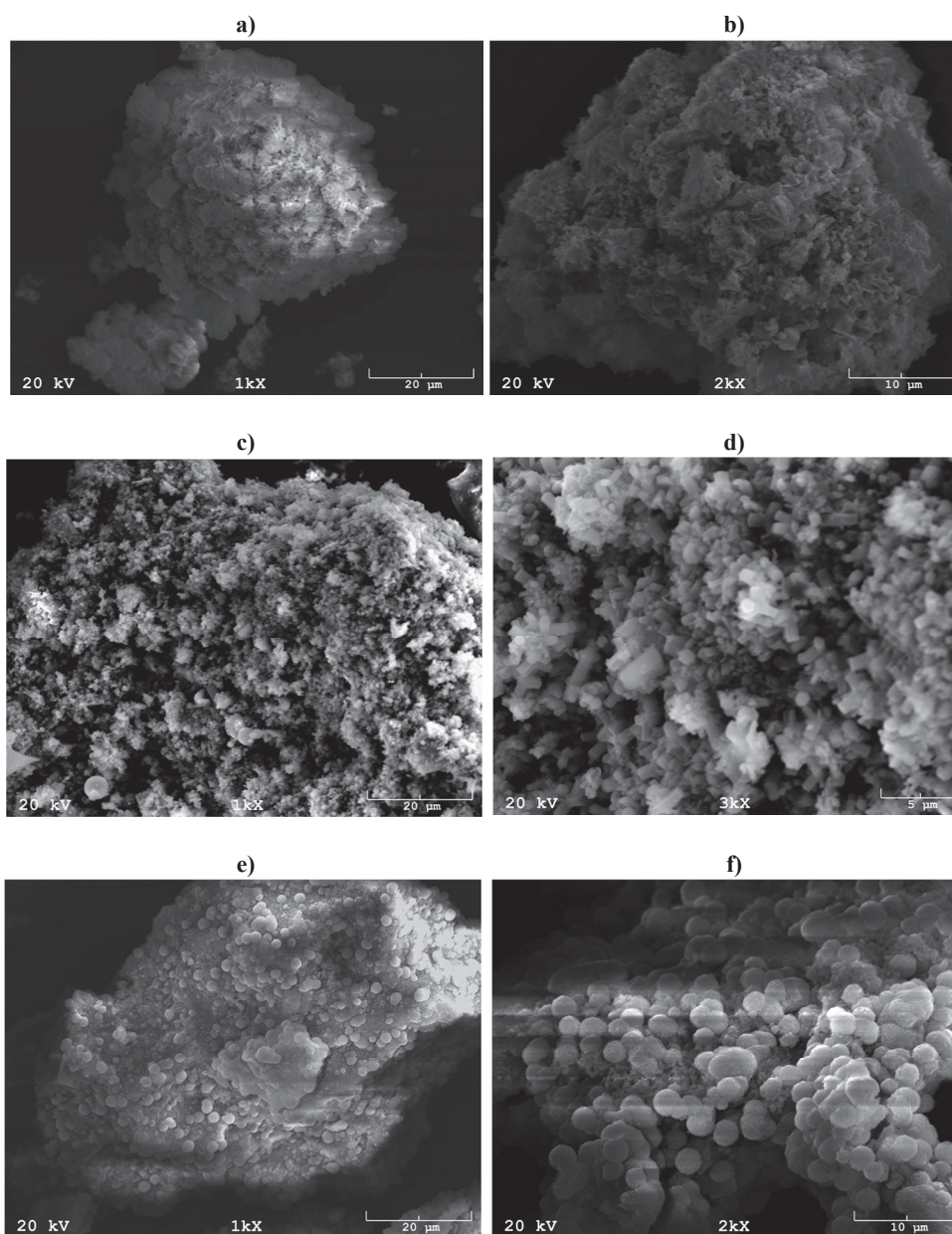


Fig. 6. X-ray diffraction pattern of samples. P2A(a): Analcime PDF No. 07-0364, Thomsonite PDF No. 46-1448, Quartz PDF No. 05-0490; and P2H (b): Analcime PDF No. 07-0364, Hydroxyapatite PDF No. 24-0033, Thomsonite PDF No. 46-1448, Quartz PDF No. 05-0490.





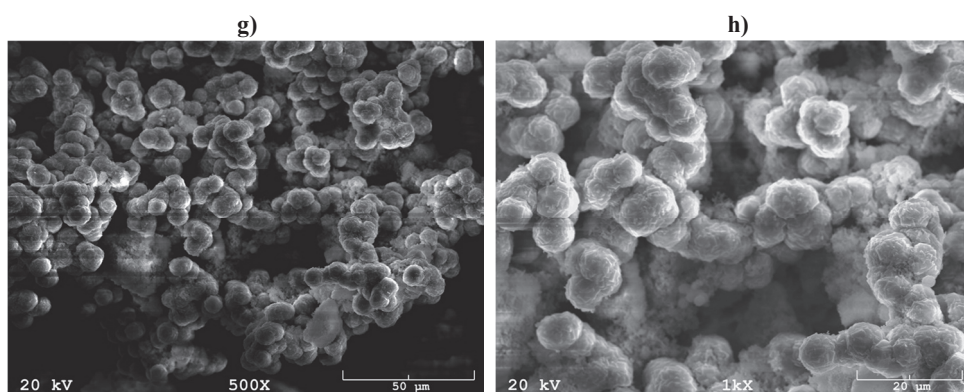
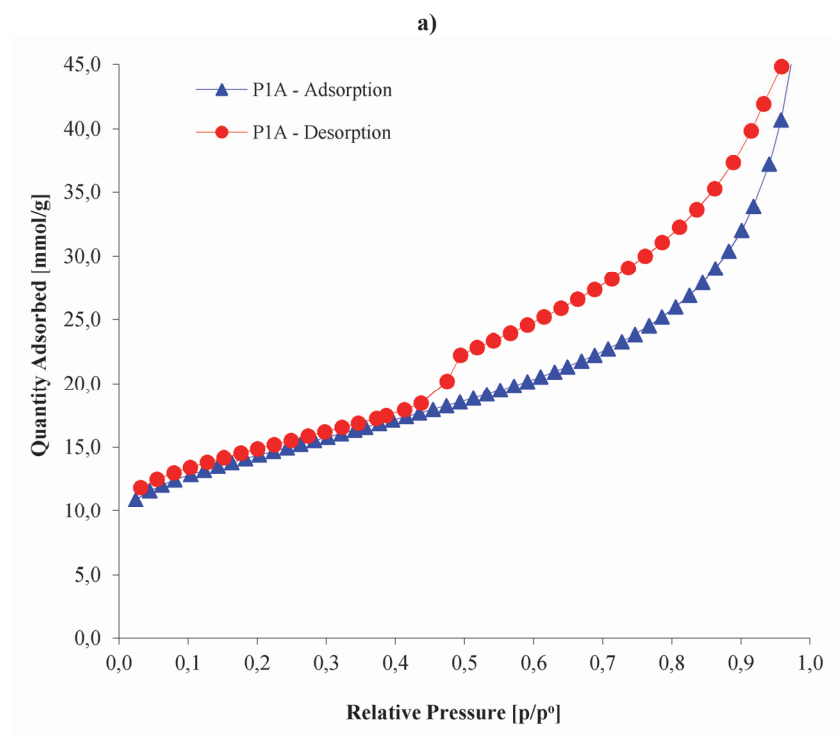
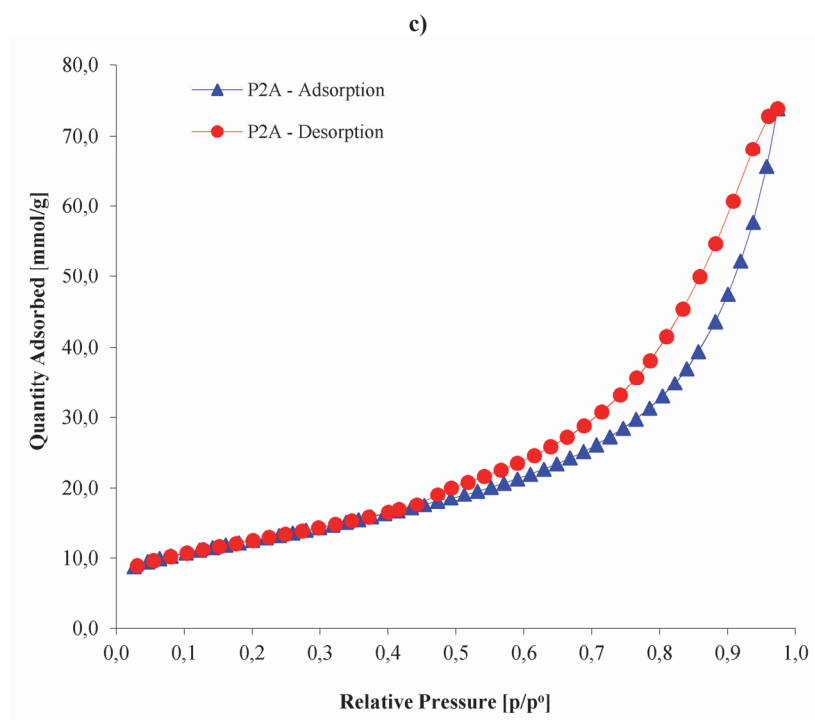
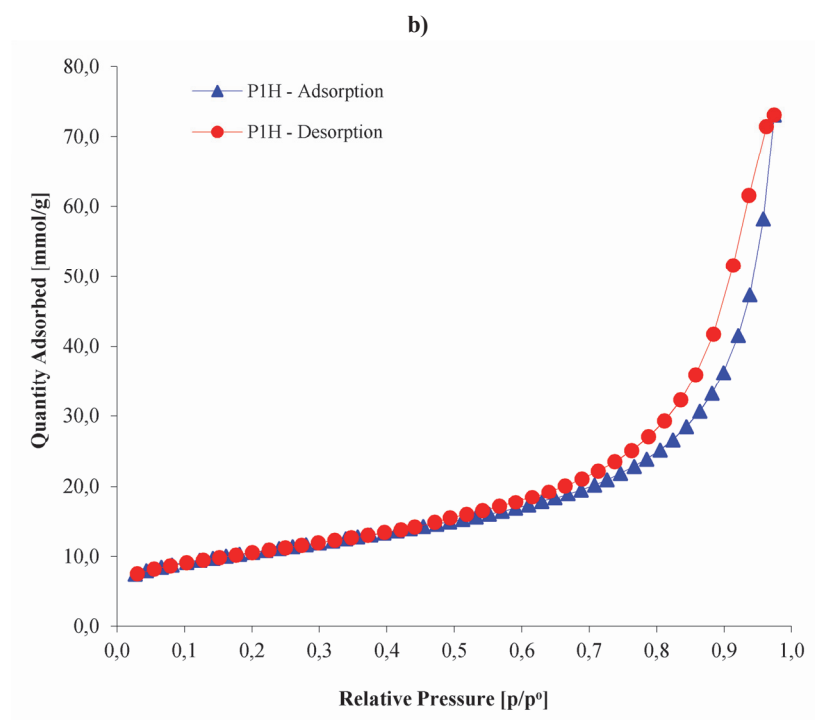


Fig. 7. SEM images of sample P1A (a, b); SEM images of sample P1H (c, d); SEM images of sample P2A (e, f); SEM images of sample P2H (g, h).







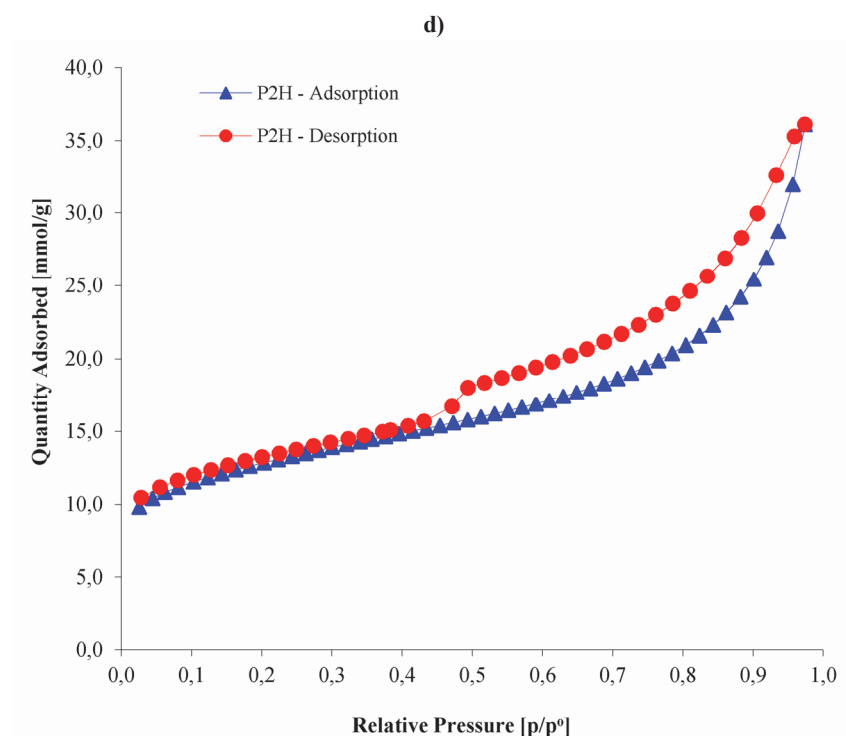


Fig. 8. BET adsorption and desorption isotherms: (a) P1A, (b) P1H, (c) P2A, (d) P2H.

Comparison of the CEC and AEC results for P1 and P1A (Table 4 and 5) as well as for P2 and P2A (Table 4 and 5) indicates that the process of autoclaving increased the cation exchange capacity of samples P1A and P2A. The cation exchange capacity of sample P1A increased three times compared to sample P1. For sample P2A, the CEC increased twenty times compared to sample P2. The AEC in both P1A and P2A samples (compared to samples P1 and P2) has not changed.

TABLE 5

Textural testing, CEC and AEC results for samples P1A, P1H, P2A and P2H.

Sample	BET [m <sup>2</sup> /g]	BJH <sup>1</sup> [m <sup>2</sup> /g]	Single point adsorption <sup>2</sup> [cm <sup>3</sup> /g]	Adsorption average <sup>3</sup> [nm]	BJH Adsorption <sup>4</sup> [nm]	CEC [meq/100g]	AEC [meq/100g]
P1A	51.49	32.43	0.071	5.50	8.18	15.18	34.39
P1H	37.09	31.60	0.113	12.20	14.47	-	-
P2A	43.30	39.38	0.114	10.60	11.85	61.39	24.80
P2H	46.26	26.76	0.056	4.80	7.61	-	-

## 4. Conclusions

This article presents the results of a comprehensive study related to the reuse of hazardous incineration plant wastes. The results of the current investigation showed obtaining different types of zeolite structures, depending on the method of synthesis. Synthesis of the alkaline waste-codes 190112 and 190114 leads to morphological changes in these materials and the creation of such zeolite forms as chabazite and thomsonite (for waste-code 190112), and also analcime and thomsonite (for waste-code 190114). Chabazite and analcime can be valuable absorbents material, for example in gas filters, membranes, and industrial water treatment. As a result of the synthesis carried out in an autoclave it is possible to obtain materials which have a BET specific surface area of above  $50 \text{ m}^2 \text{ g}^{-1}$  (for waste-code 190112) and above  $40 \text{ m}^2 \text{ g}^{-1}$  (for waste-code 190114).

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